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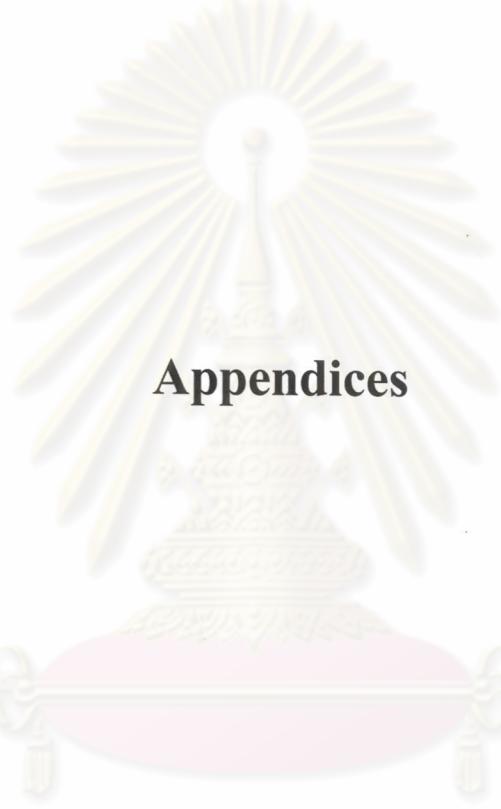
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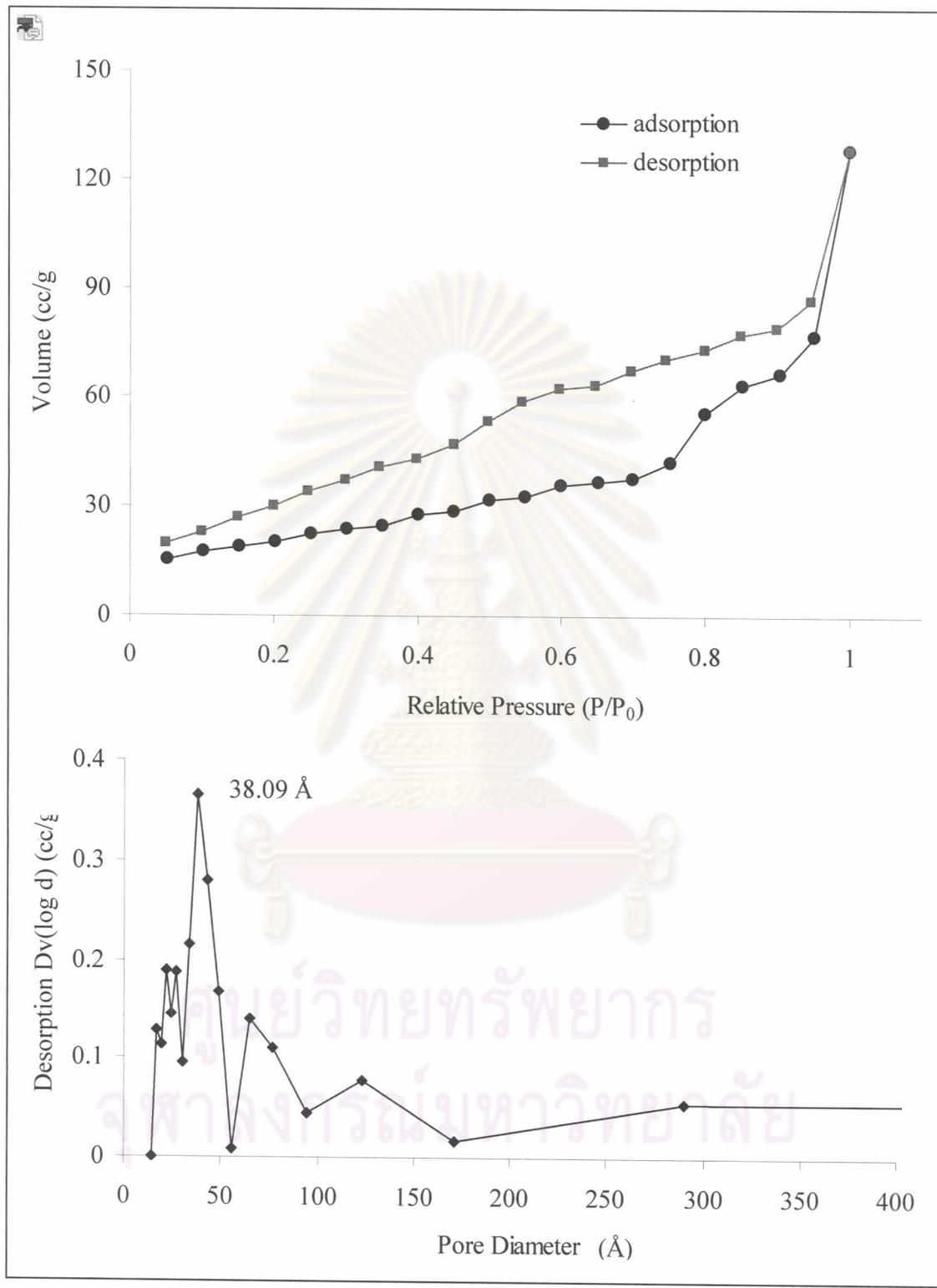
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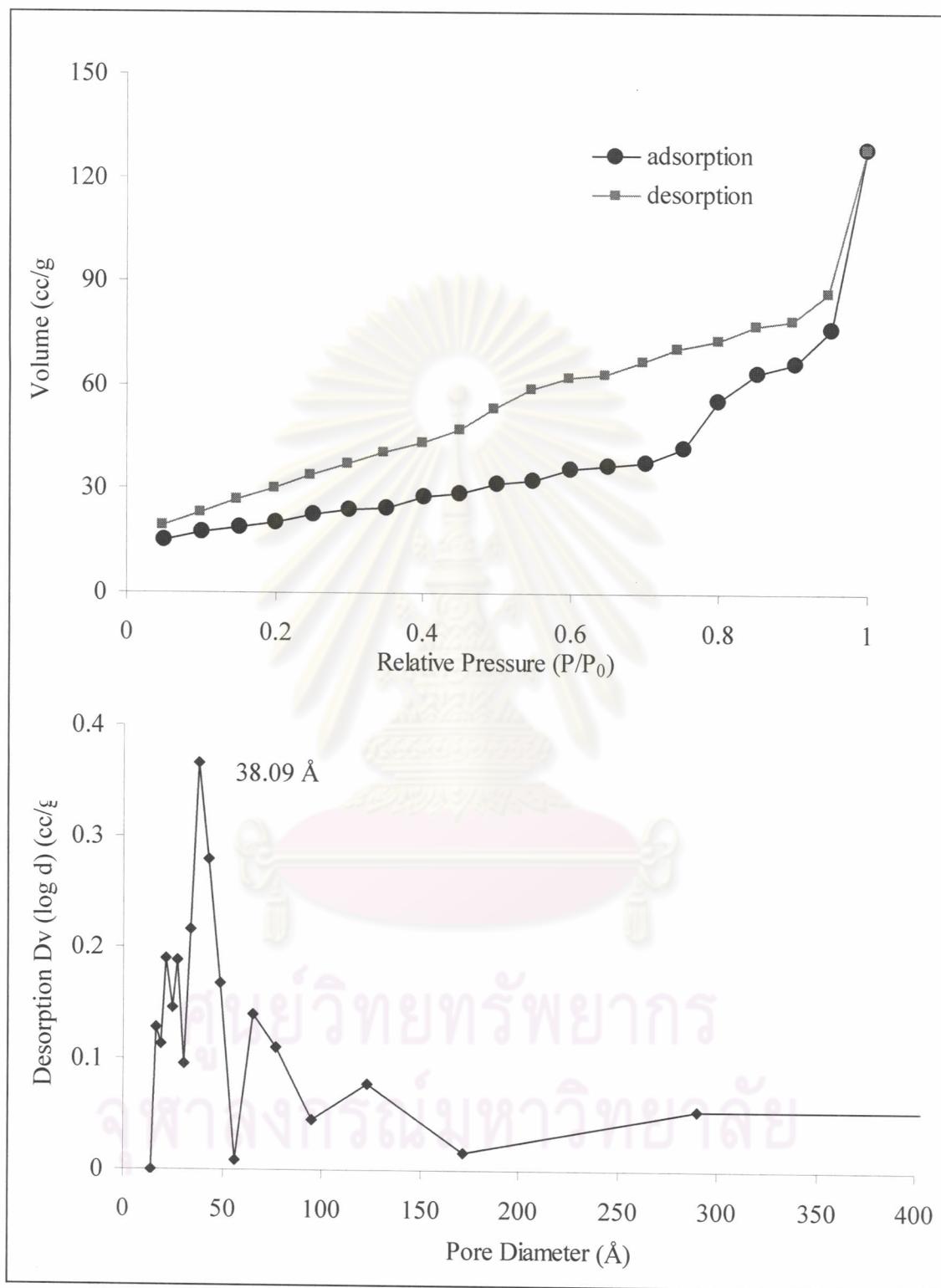


## **Appendices**

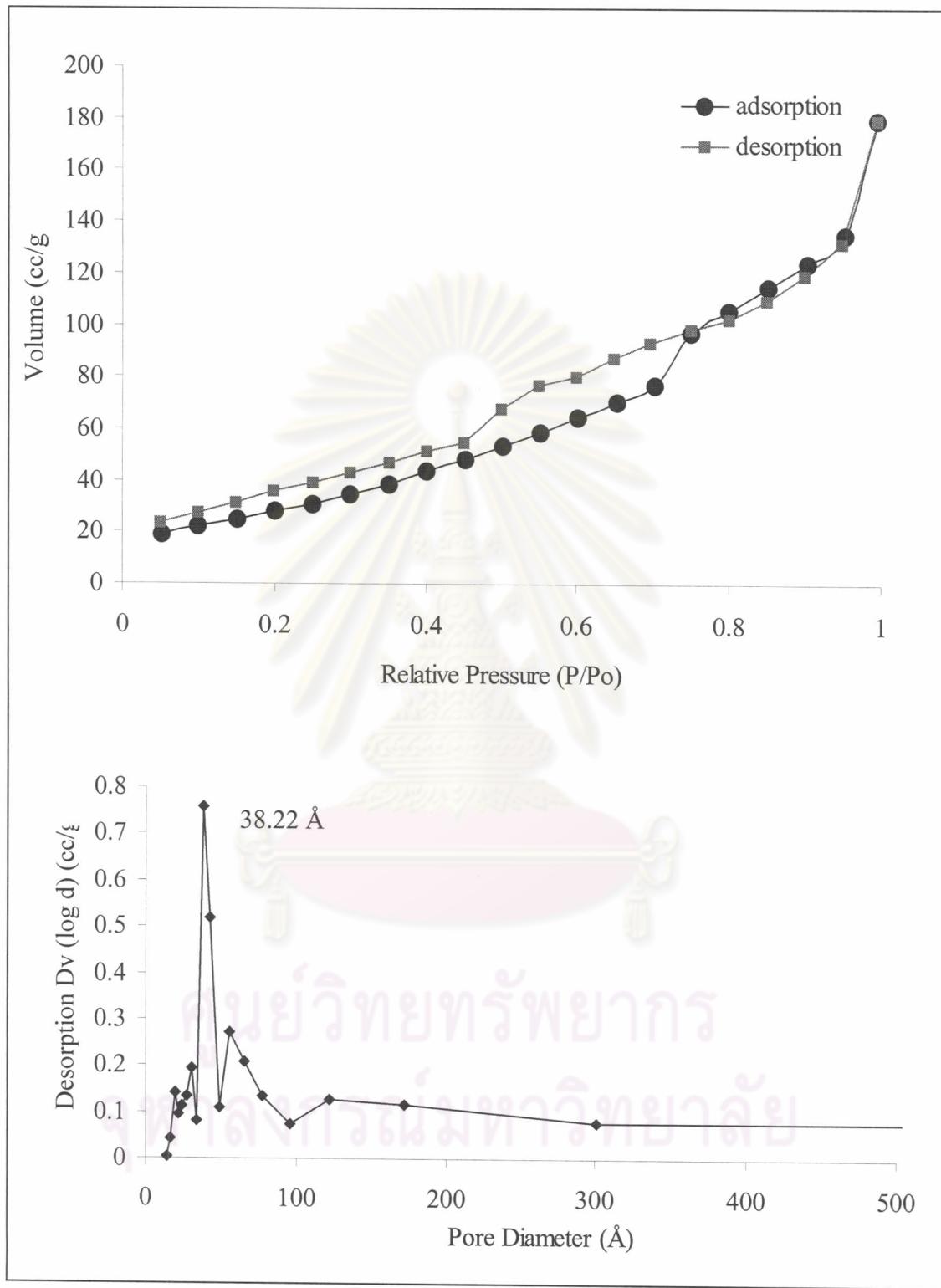
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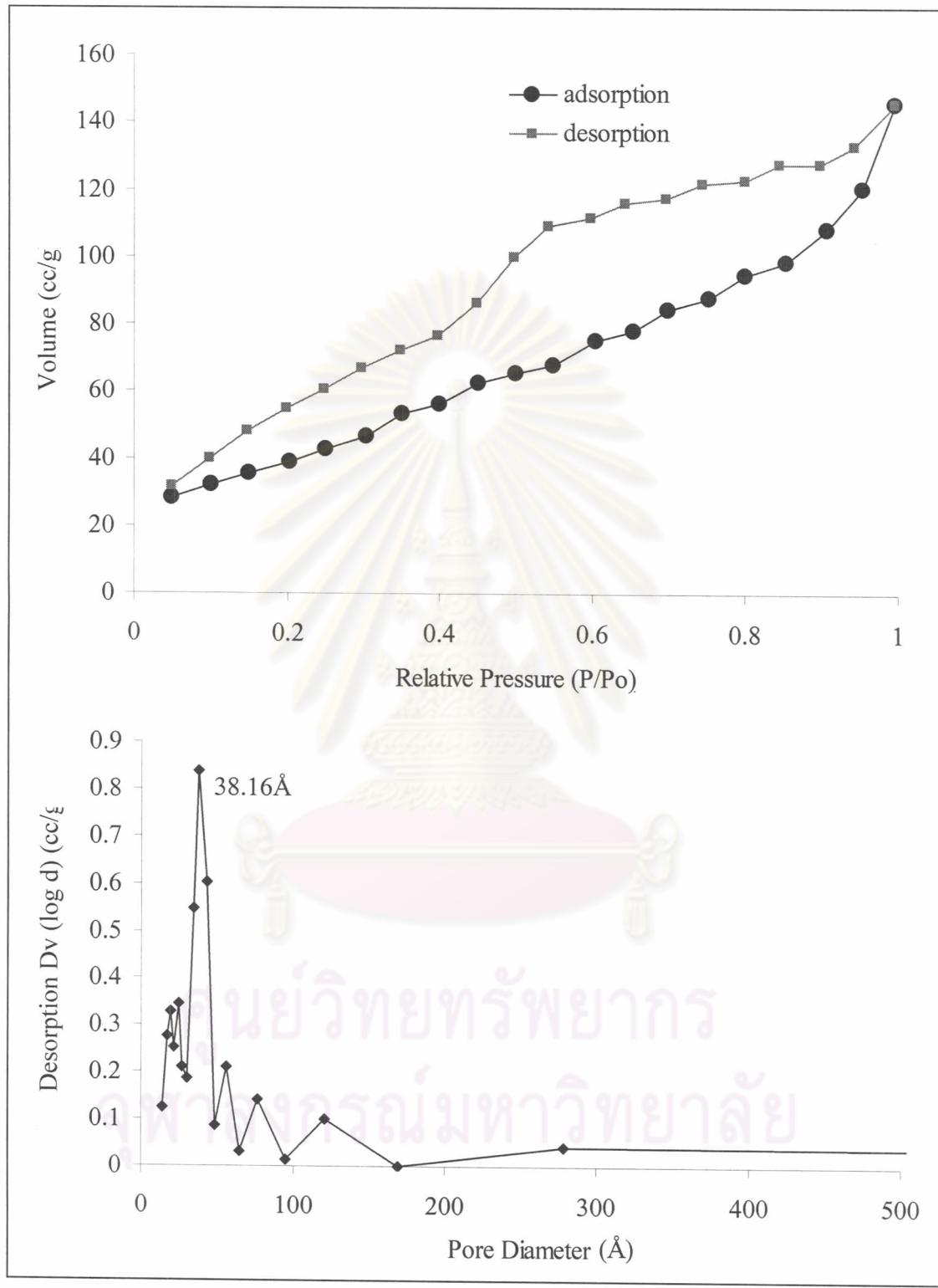
**Figure A-1** N<sub>2</sub> adsorption-desorption isotherm and pore sizes distribution of hectorite.



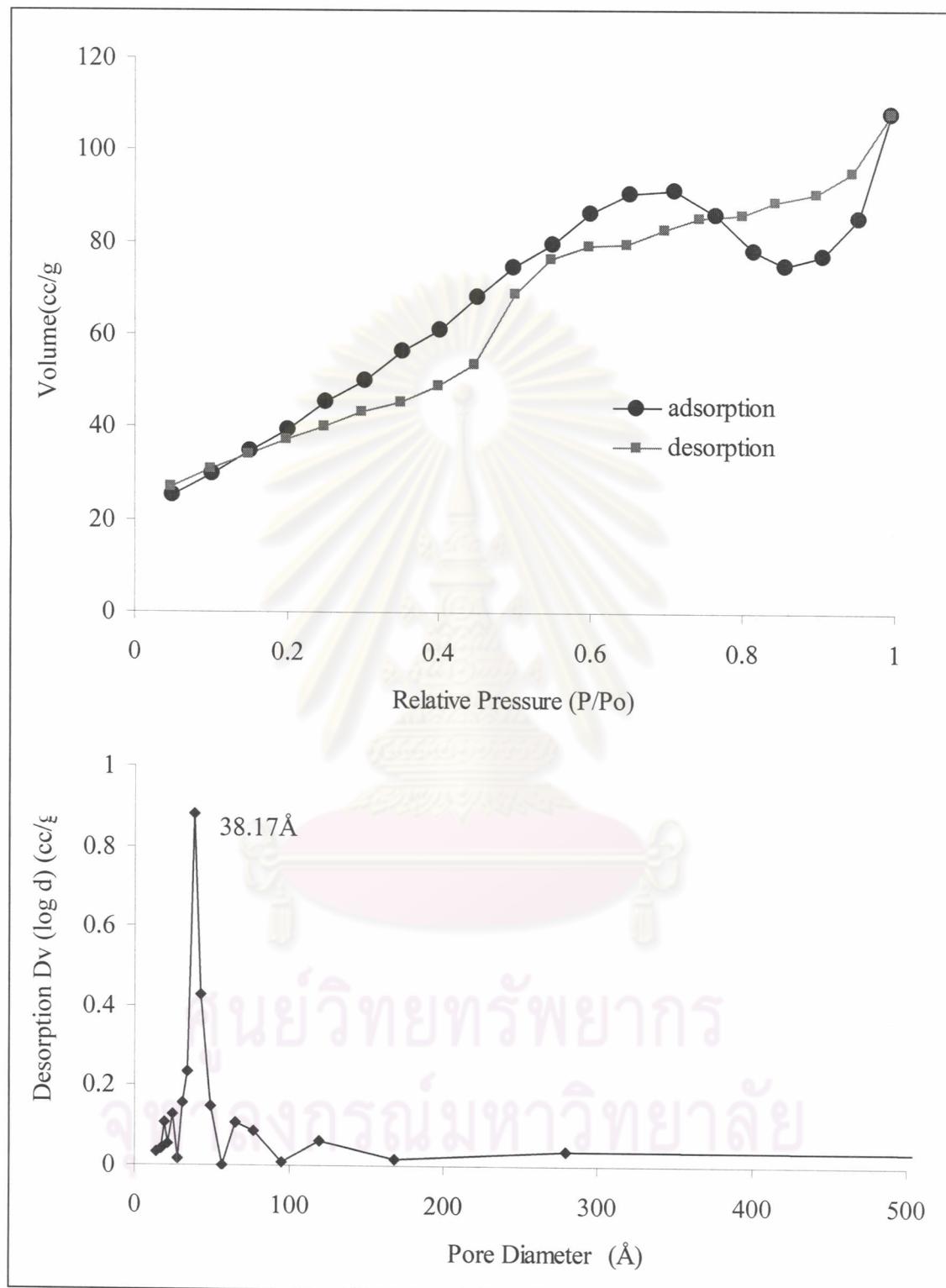
**Figure A-2** N<sub>2</sub> adsorption-desorption isotherm and pore sizes distribution of HFe<sub>0.5</sub>.



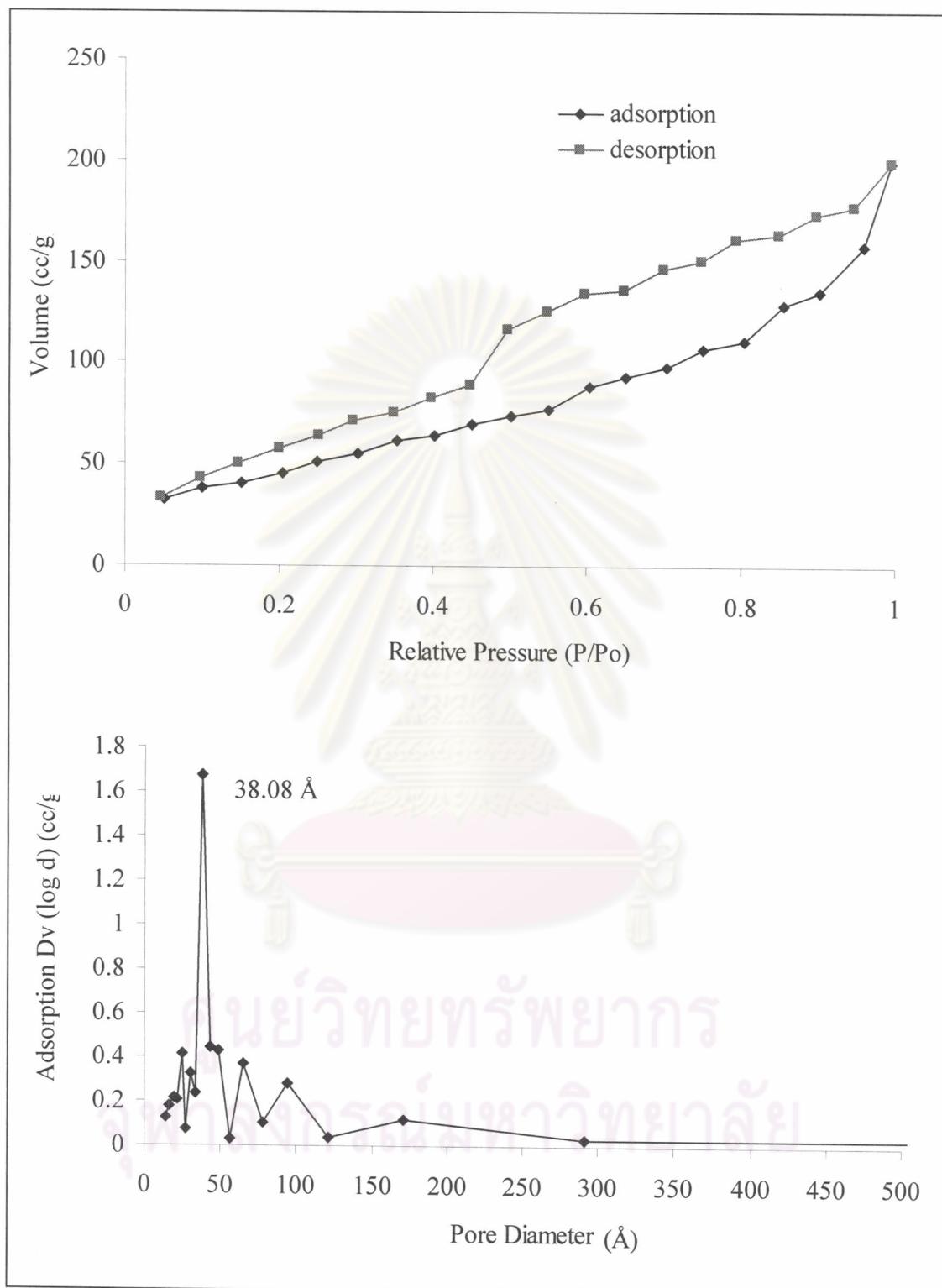
**Figure A-3** N<sub>2</sub> adsorption-desorption isotherm and pore sizes distribution of HFe<sub>1</sub>.



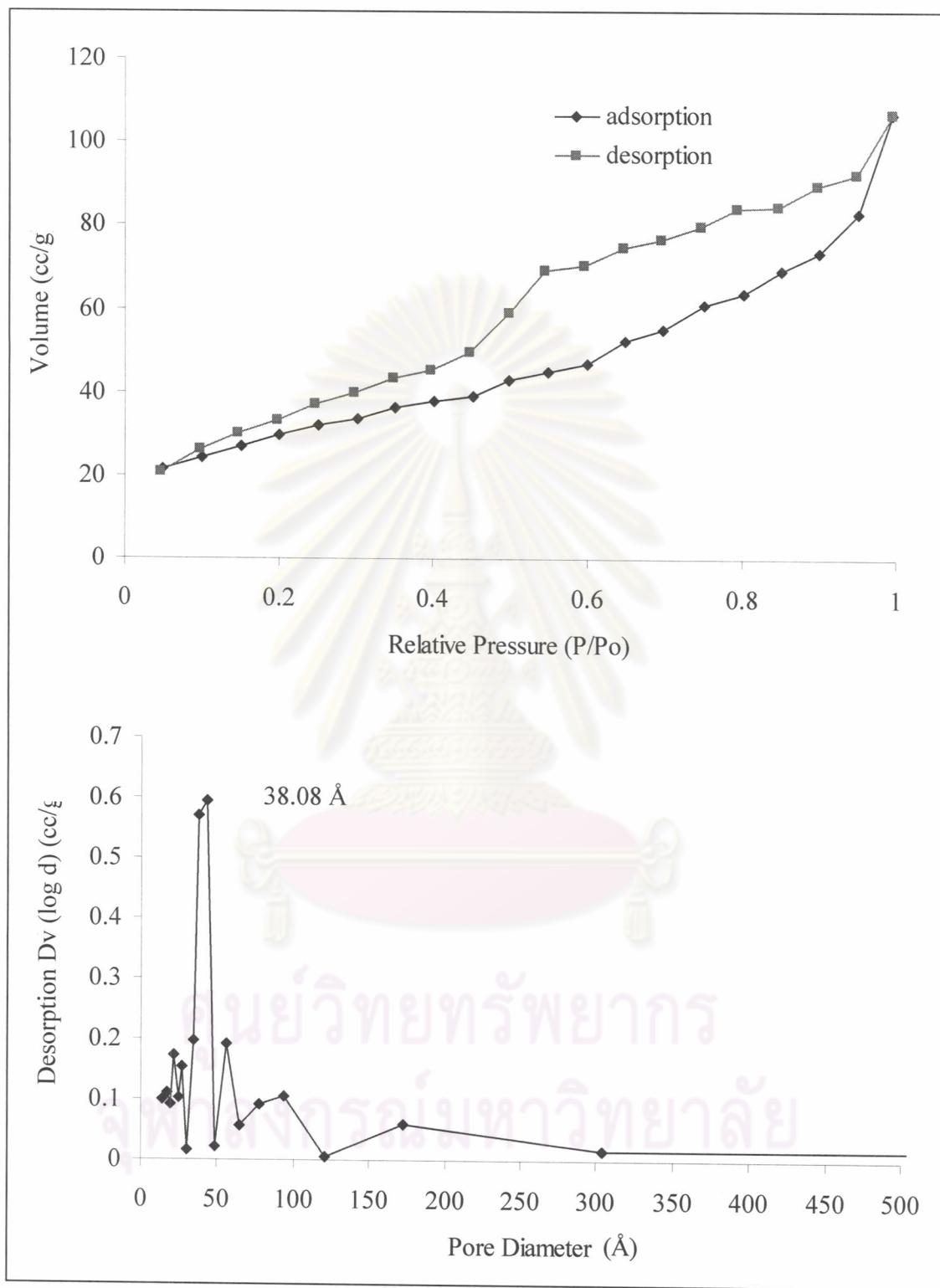
**Figure A-4** N<sub>2</sub> adsorption-desorption isotherm and pore sizes distribution of HFe<sub>5</sub>.



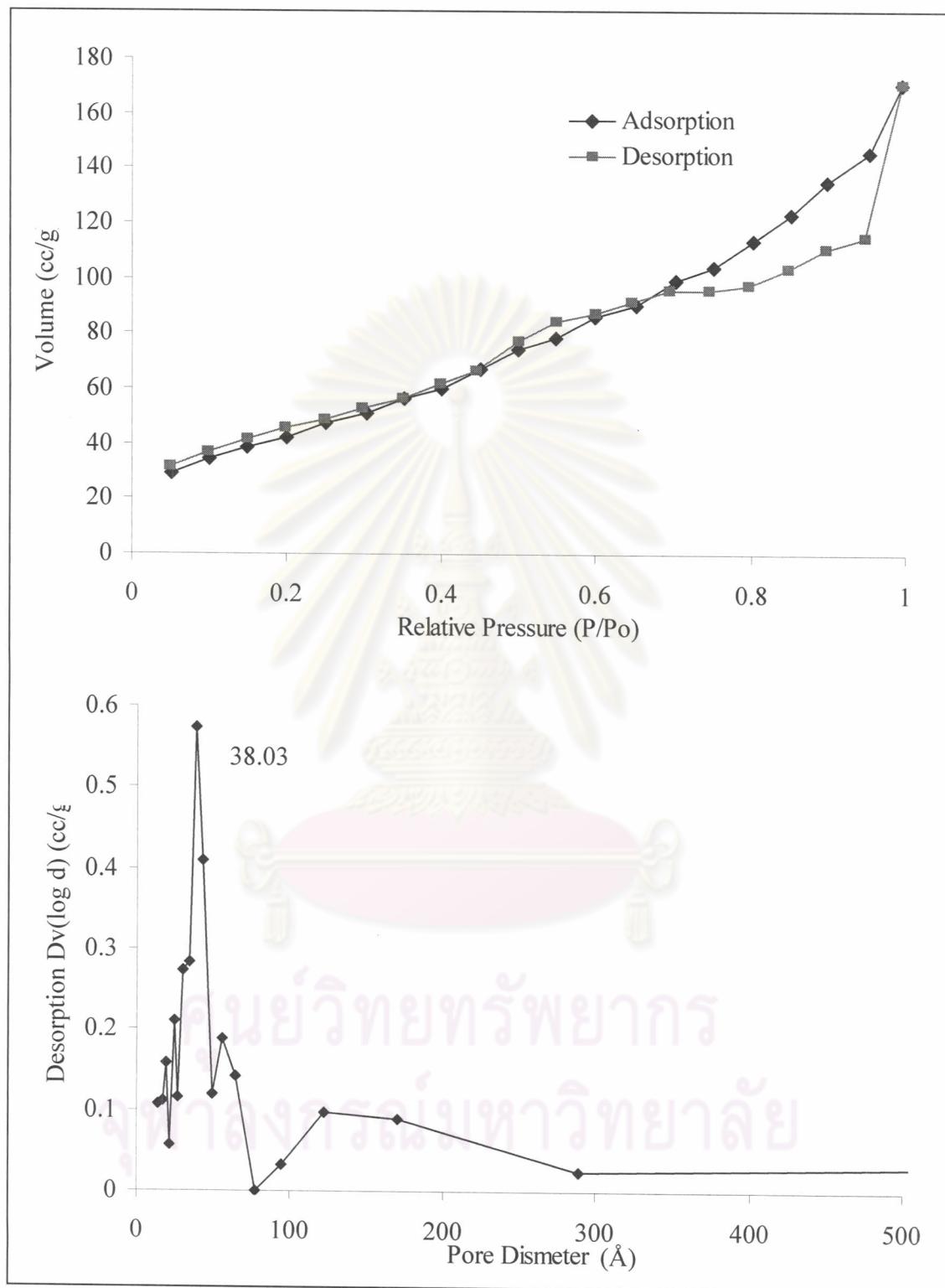
**Figure A-5** N<sub>2</sub> adsorption-desorption isotherm and pore sizes distribution of HFe<sub>10</sub>.



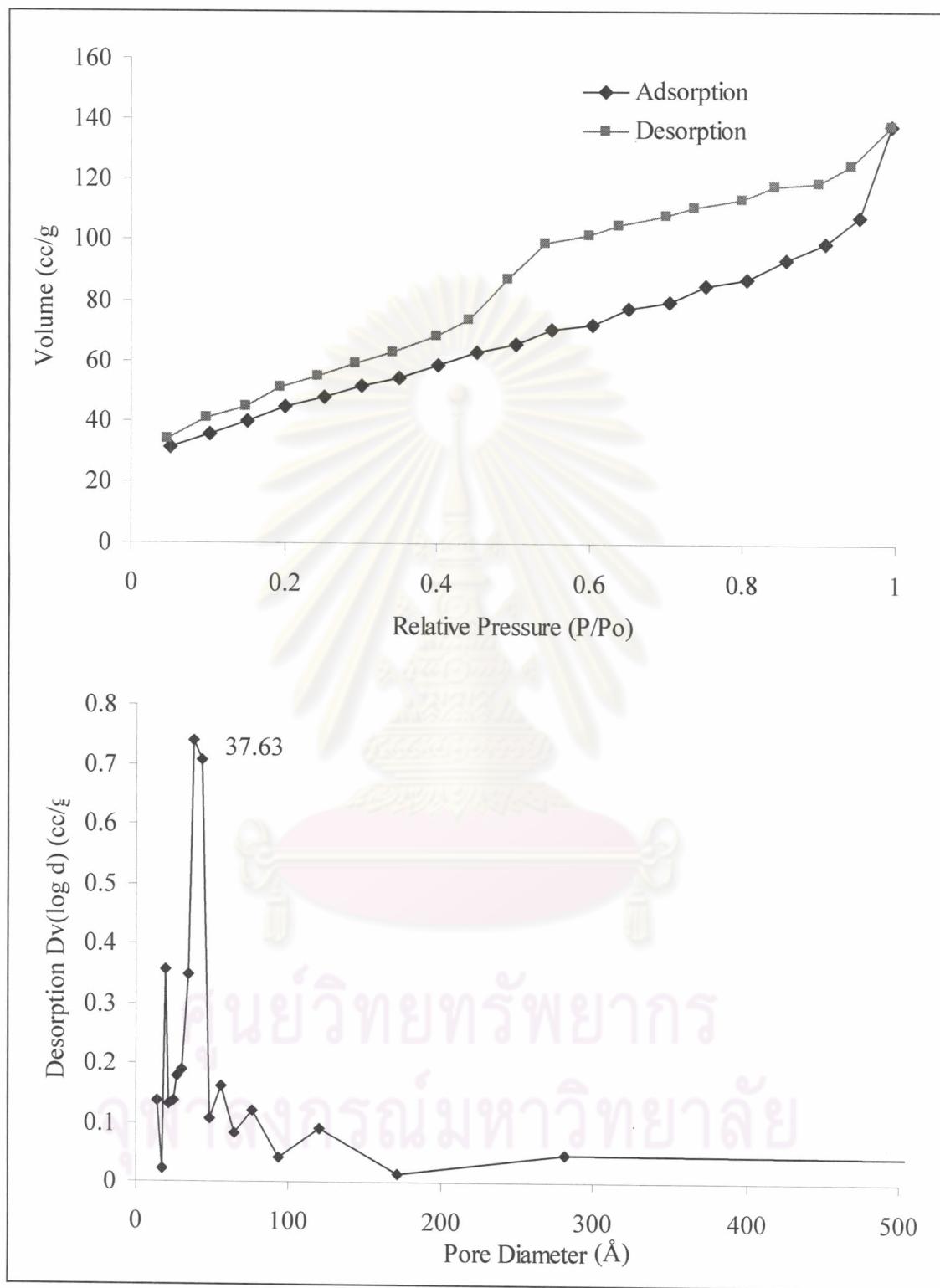
**Figure A-6** N<sub>2</sub> adsorption-desorption isotherm and pore sizes distribution of HFe<sub>20</sub>.



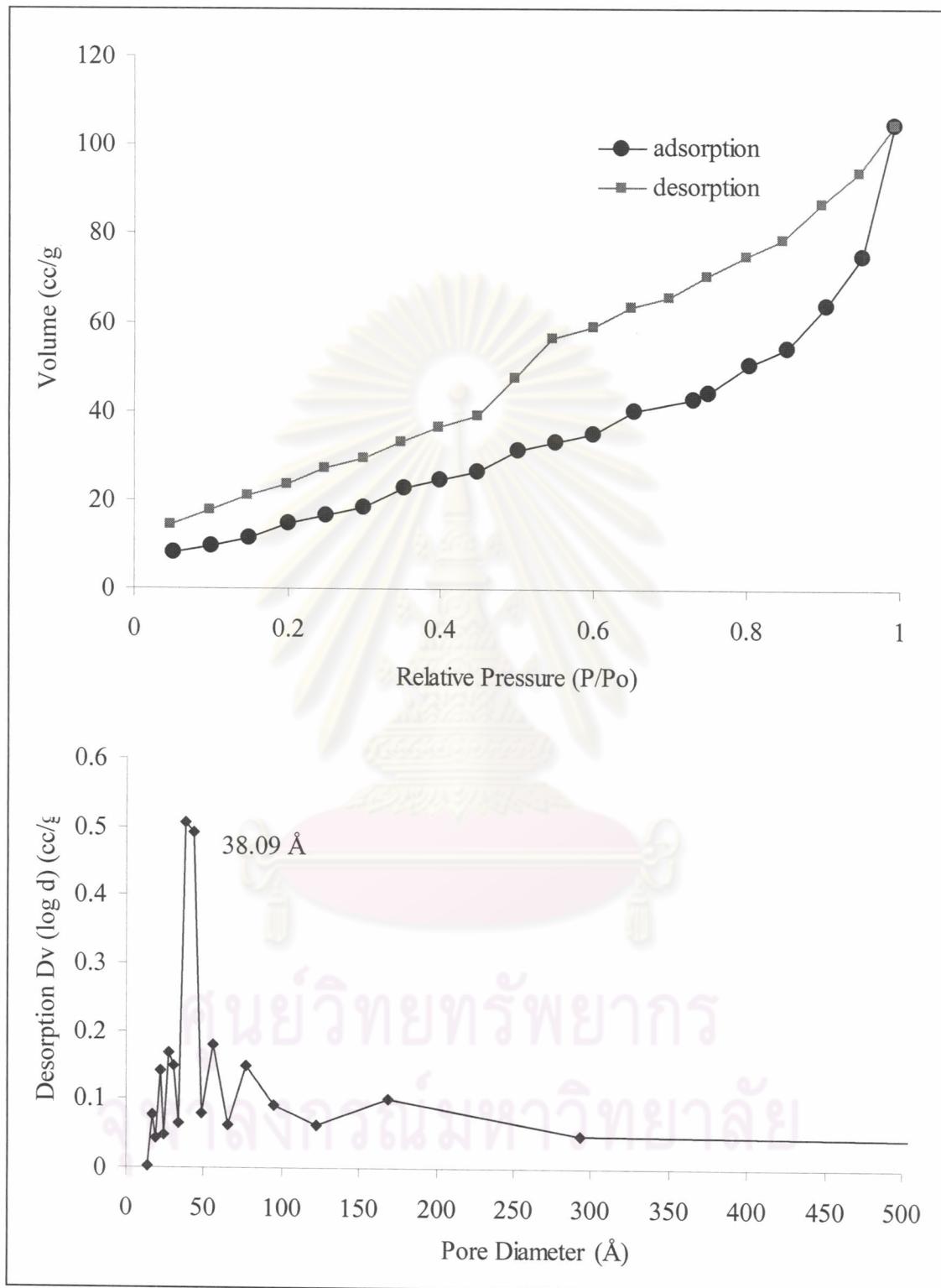
**Figure A-7** N<sub>2</sub> adsorption-desorption isotherm and pore sizes distribution of HFe<sub>240</sub>.



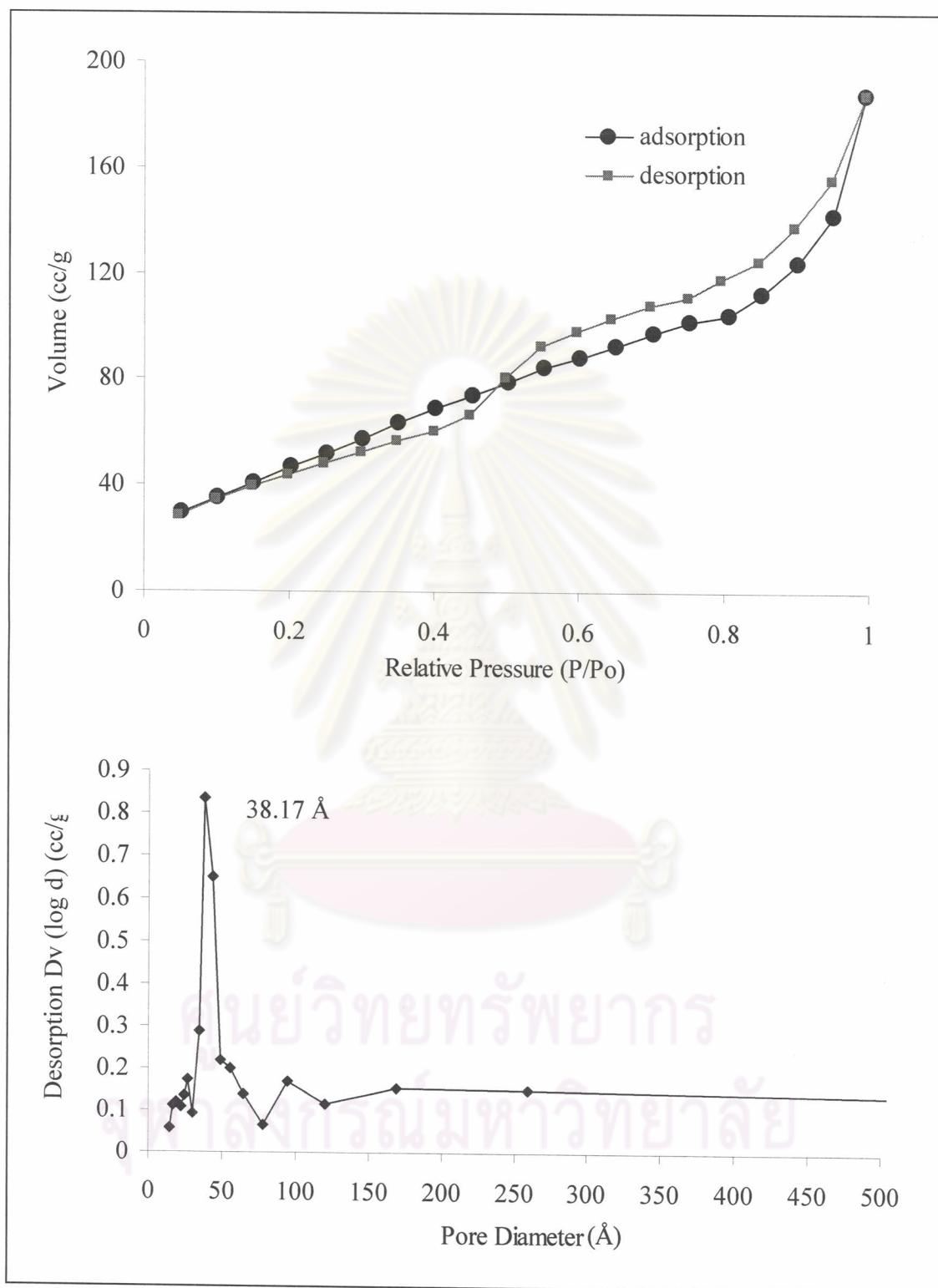
**Figure A-8**  $N_2$  adsorption-desorption isotherm and pore sizes distribution of  $HFe_{10}Ga_1$ .



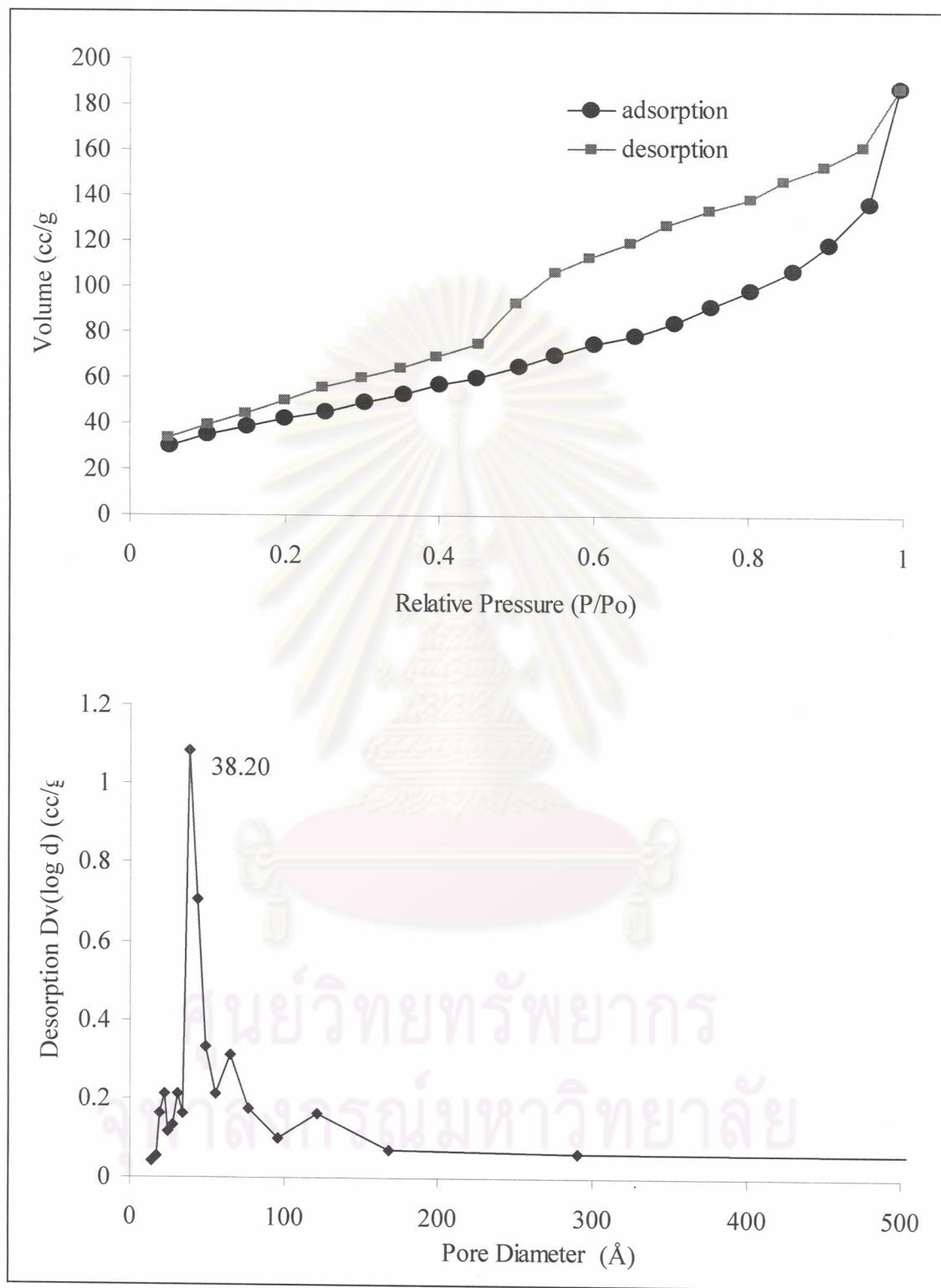
**Figure A-9** N<sub>2</sub> adsorption-desorption isotherm and pore sizes distribution of HFe<sub>10</sub>Ga<sub>1</sub>I.



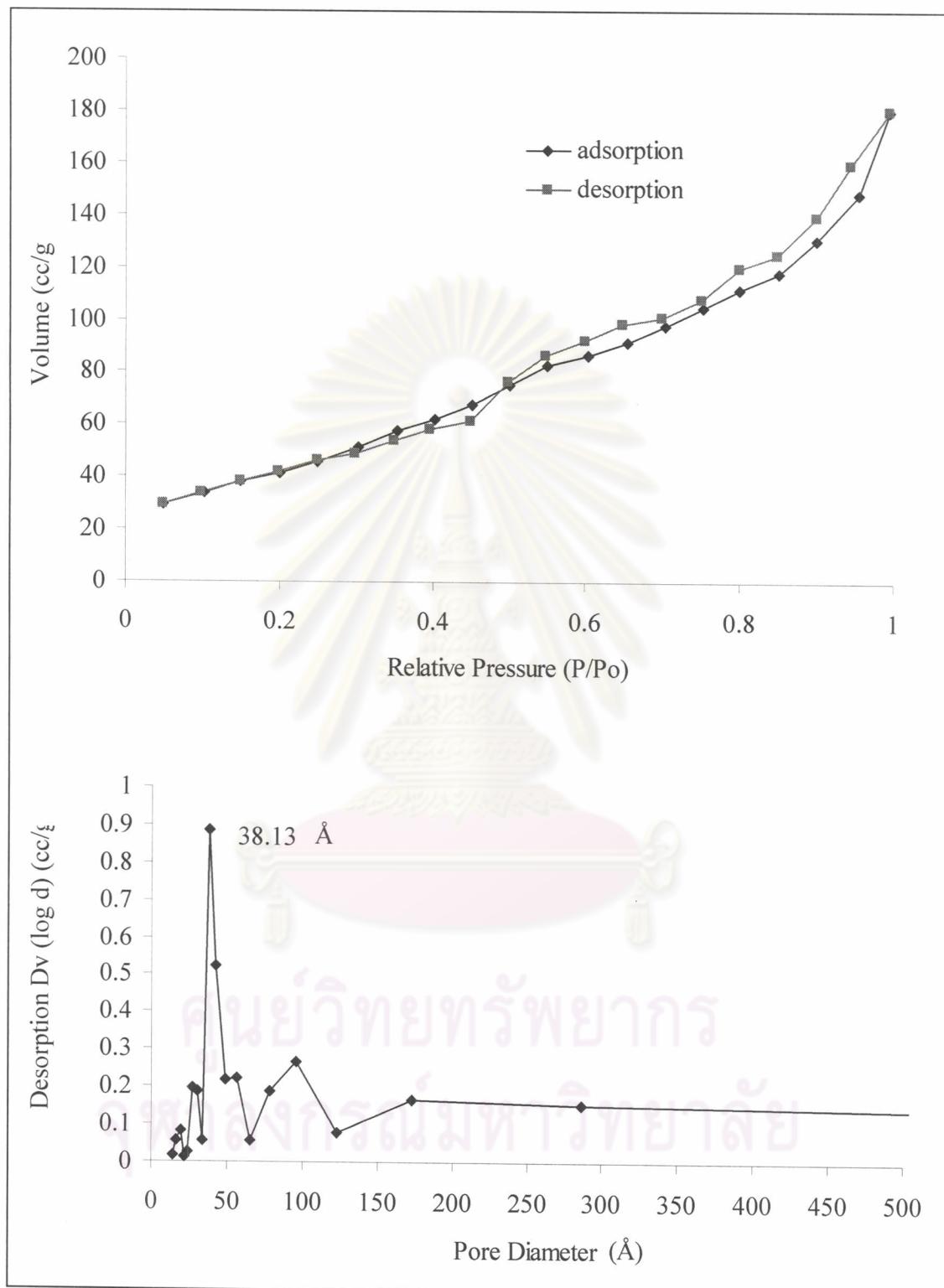
**Figure A-10** N<sub>2</sub> adsorption-desorption isotherm and pore sizes distribution of bentonite.



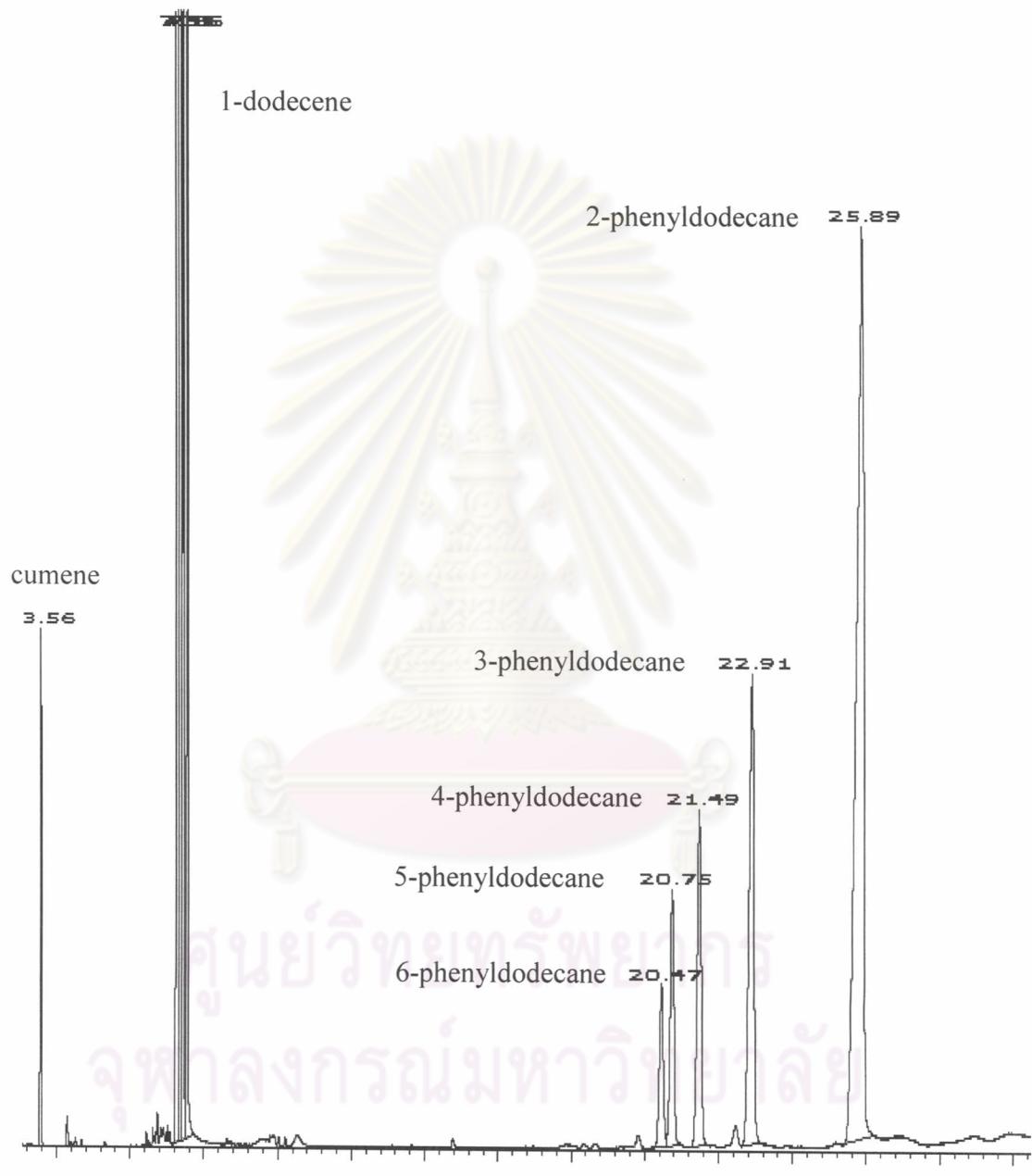
**Figure A-11** N<sub>2</sub> adsorption-desorption isotherm and pore sizes distribution of BFe<sub>5</sub>.



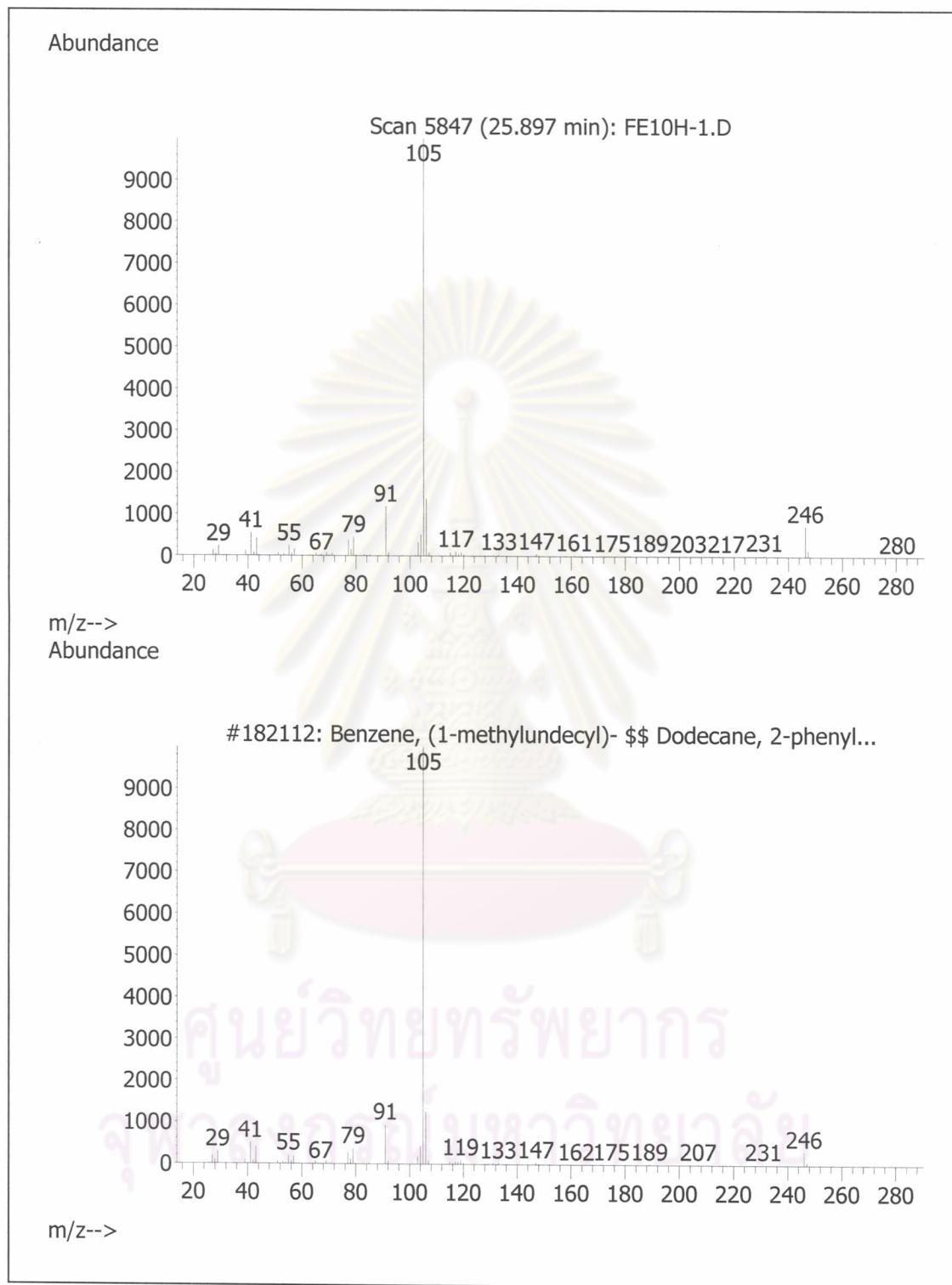
**Figure A-12**  $N_2$  adsorption-desorption isotherm and pore sizes distribution of  $BFe_{10}$ .



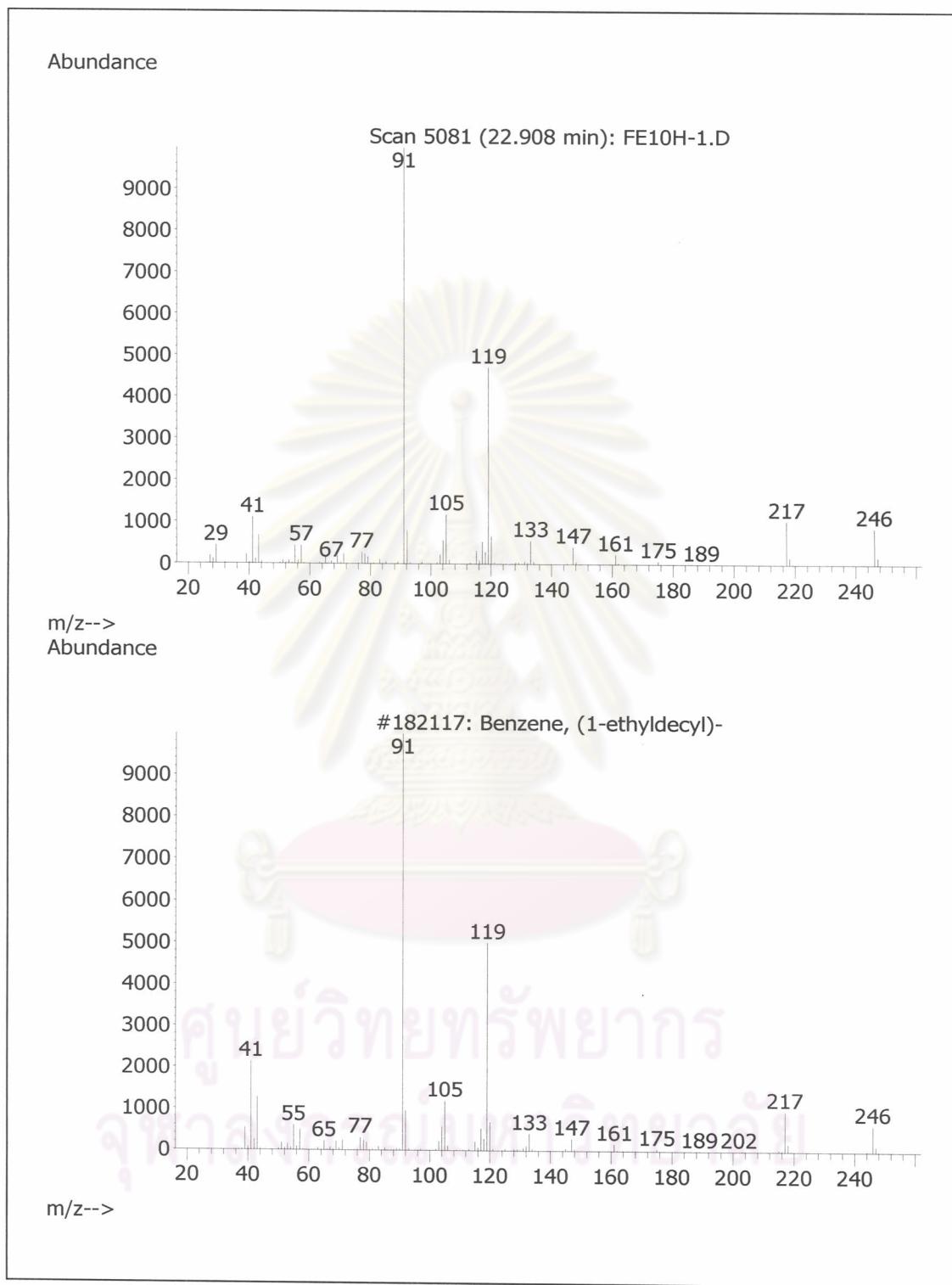
**Figure A-13**  $N_2$  adsorption-desorption isotherm and pore sizes distribution of  $BFe_{20}$ .



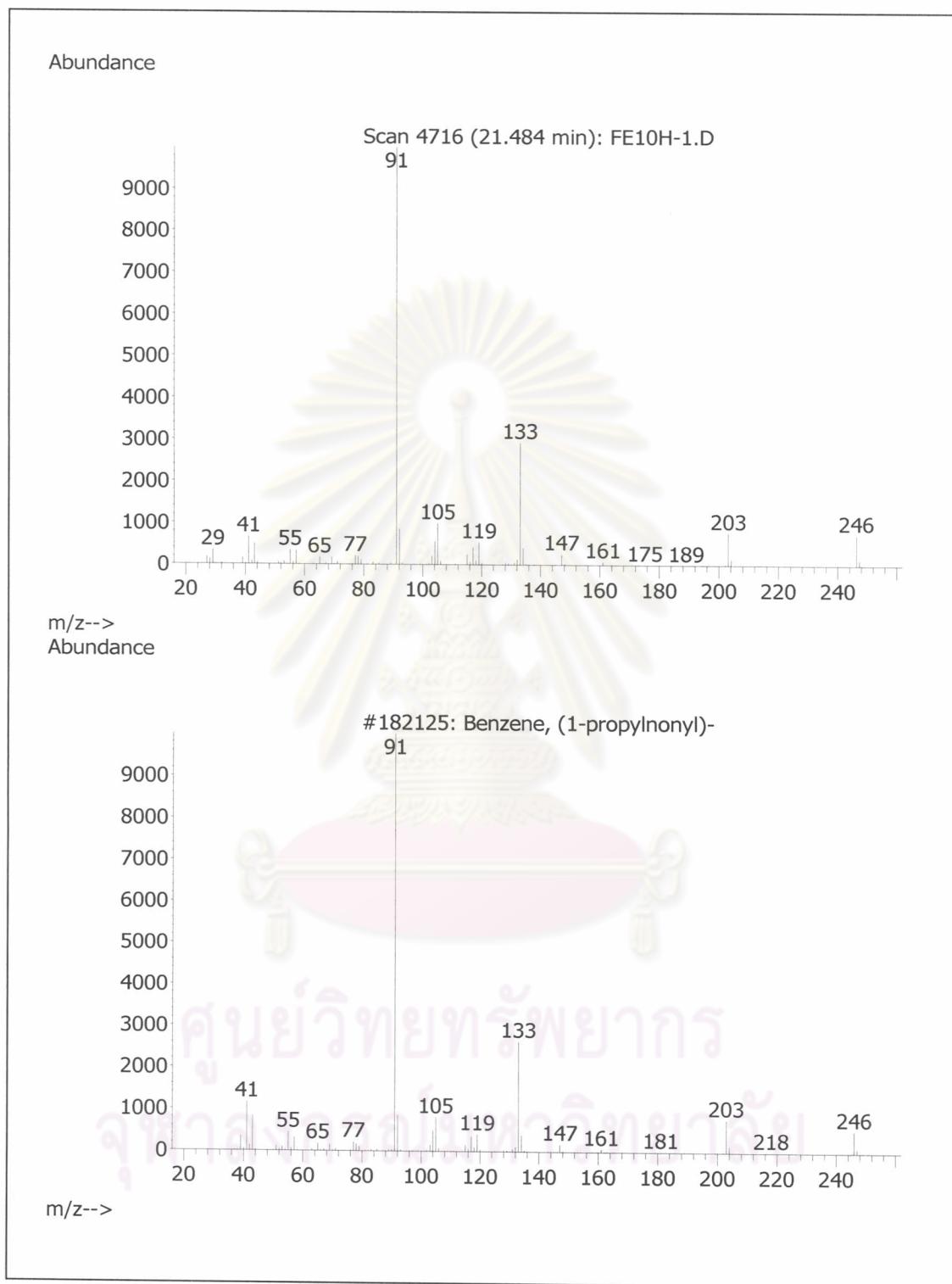
**Figure A-14** A gas chromatogram of liquid product from alkylation of benzene with 1-dodecene.



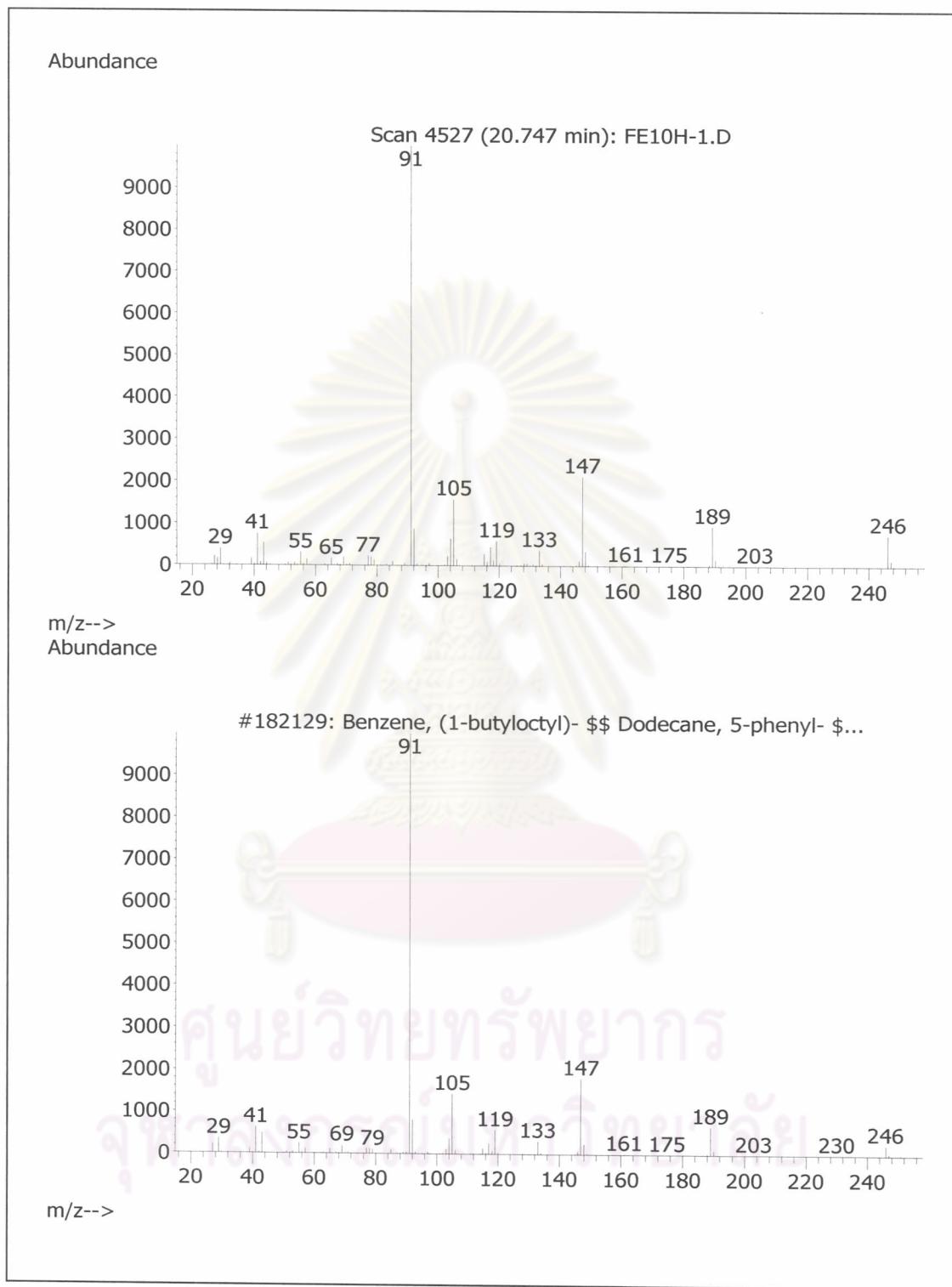
**Figure A-15** Mass spectrum of liquid product at  $t_R = 25.897\text{s}$  (top) and 2-phenyldodecane in library (bottom).



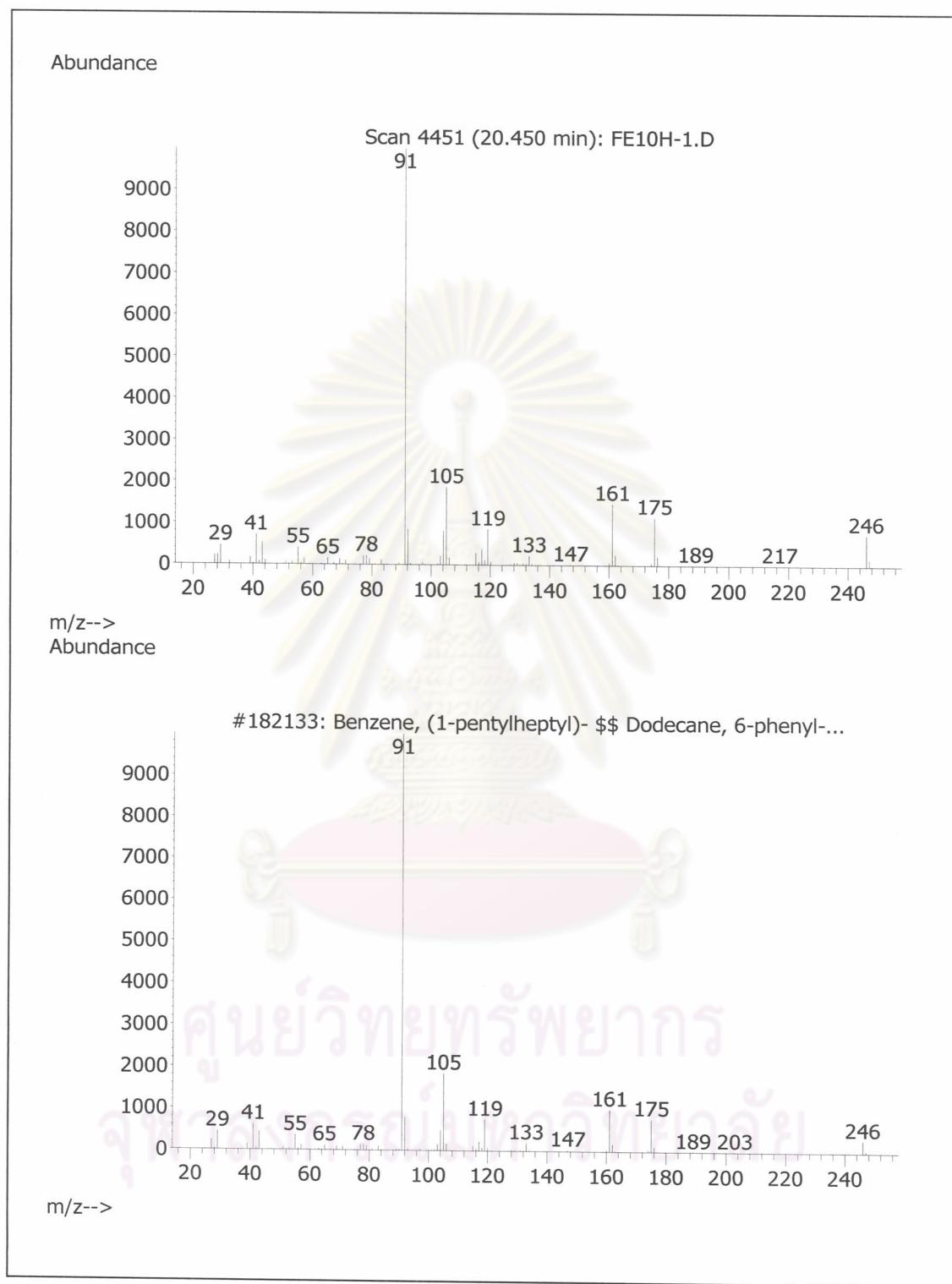
**Figure A-16** Mass spectrum of liquid product at  $t_R = 22.908\text{s}$  (top) and 3-phenyldodecane in library (bottom).



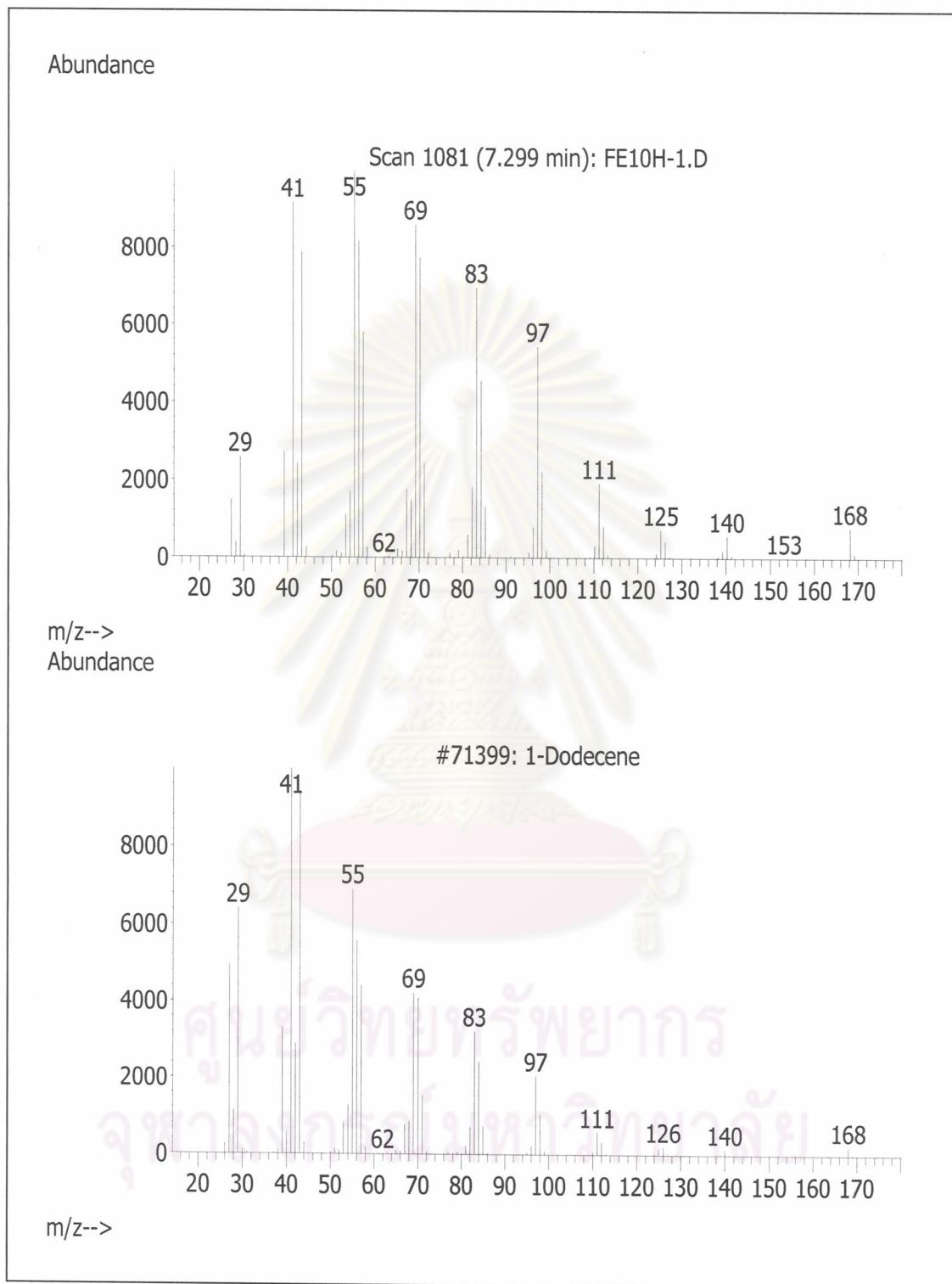
**Figure A-17** Mass spectrum of liquid product at  $t_R = 21.484\text{s}$  (top) and 4-phenyldodecane in library (bottom).



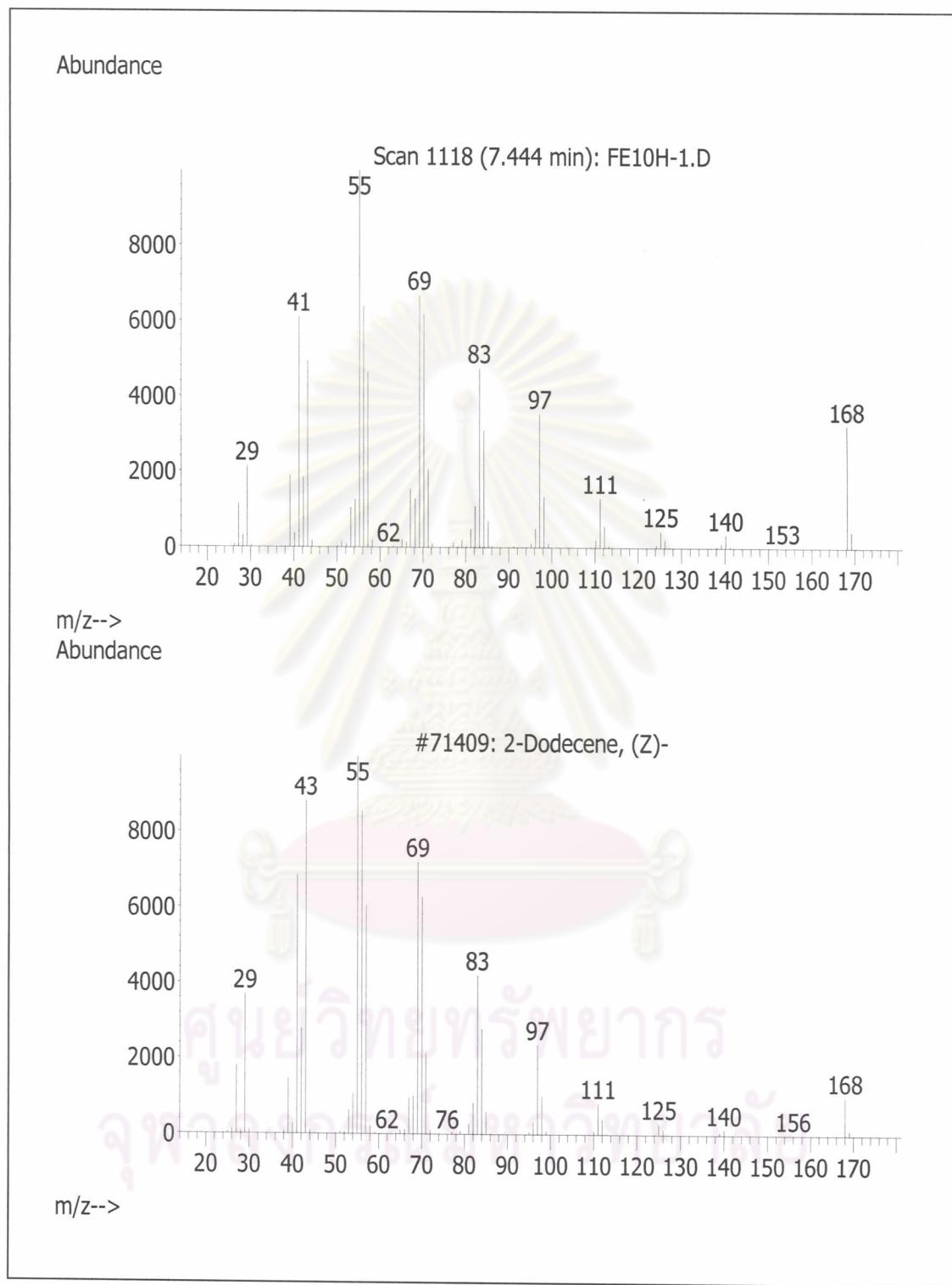
**Figure A-18** Mass spectrum of liquid product at  $t_R = 20.747\text{s}$  (top) and 5-phenyldodecane in library (bottom).



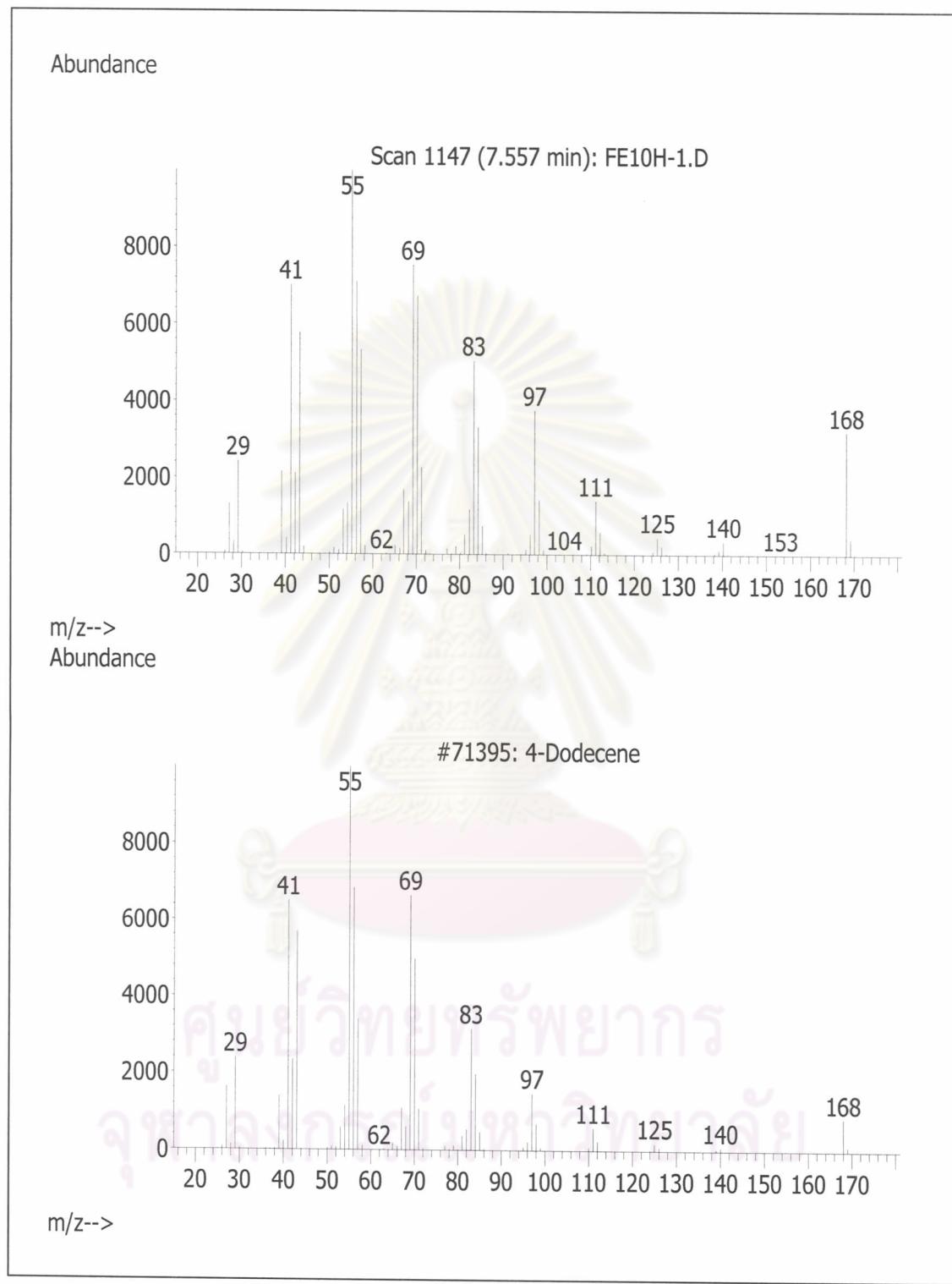
**Figure A-19** Mass spectrum of liquid product at  $t_R = 20.450\text{s}$  (top) and 6-phenyldodecane in library (bottom).



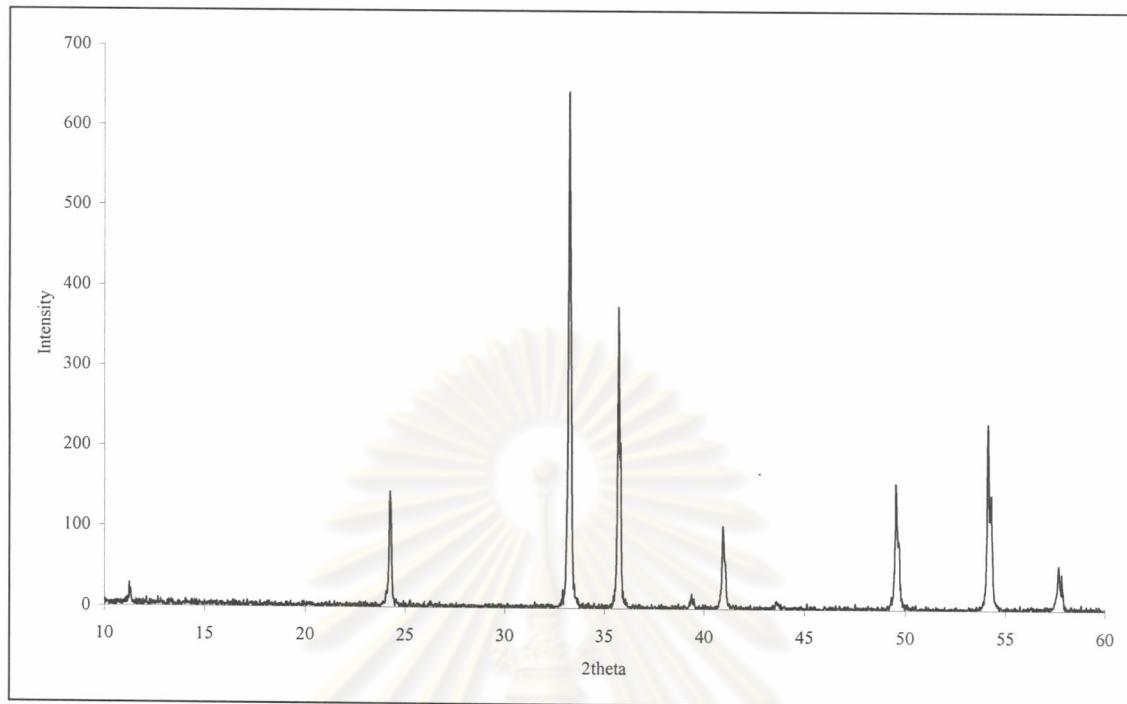
**Figure A-20** Mass spectrum of liquid product at  $t_R = 7.30\text{s}$  (top) and 1-dodecene in library (bottom).



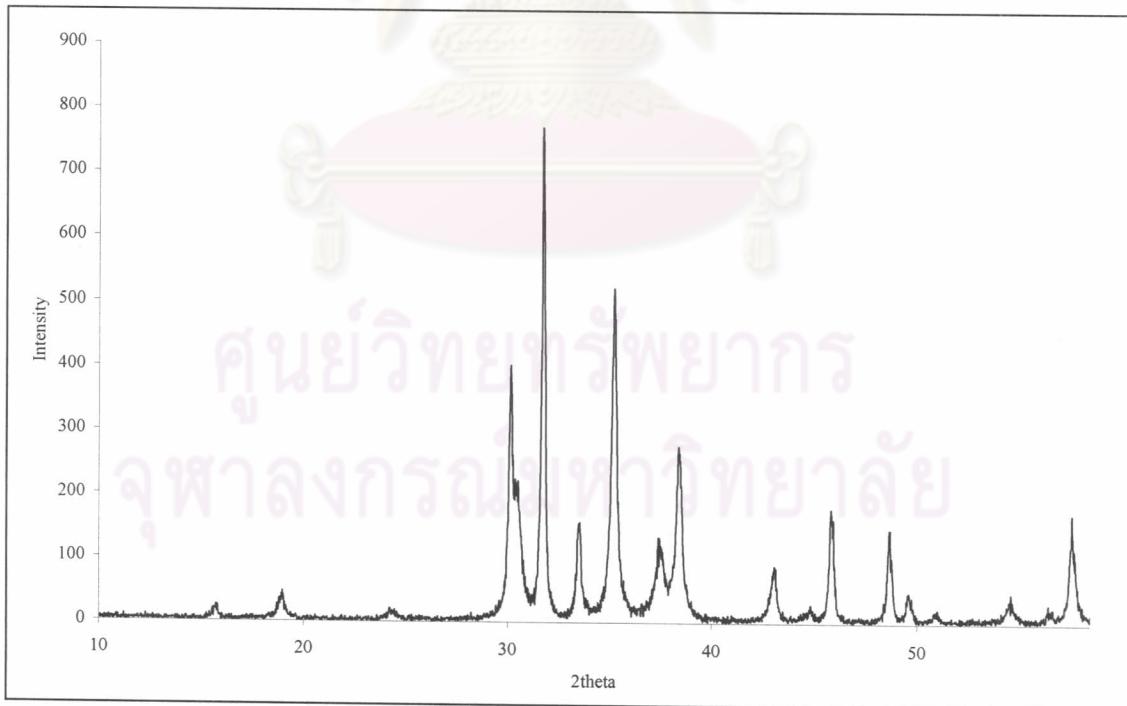
**Figure A-21** Mass spectrum of liquid product at  $t_R = 7.44\text{s}$  (top) and 2-dodecane in library (bottom).



**Figure A-22** Mass spectrum of liquid product at  $t_R = 7.56\text{s}$  (top) and 4-dodecene in library (bottom).



**Figure A-23** XRD pattern of iron oxide ( $\text{Fe}_2\text{O}_3$ , hematite)



**Figure A-24** XRD pattern of gallium oxide ( $\text{Ga}_2\text{O}_3$ ).

**Table A-1** The  $d_{001}$  spacing of hectorites treated with 2M, 5M NaCl and 5M NaOH

Number of exchange	$d_{001}$ (Å)		
	2M NaCl	5M NaCl	5M NaOH
0(Hectorite)	12.29		
1	12.38	13.00	13.11
3	13.33	14.20	14.34
4	-	14.48	15.01
5	12.62	14.67	14.96
6	-	14.40	14.82

**Calculation of %Conversion**

$$\% \text{conversion} = \frac{A_{\text{before}} - A_{\text{after}}}{A_{\text{before}}} \times 100$$

$A_{\text{before}}$  = area of 1-dodecene peak before reaction

$A_{\text{after}}$  = area of 1-dodecene peak after reaction

**Calculation of %Selectivity**

$$\% \text{selectivity} = \frac{\text{mol}_{\text{interest}}}{\text{mol}_{\text{total products}}} \times 100$$

$\text{mol}_{\text{interest}}$  = mol of interested product

$\text{mol}_{\text{total products}}$  = total mol of products

$$\text{mol}_{\text{interest}} = \frac{A_{\text{interest}} \times C_{\text{std}} \times \text{mol}_{\text{std}}}{A_{\text{std}} \times C_{\text{interest}}}$$

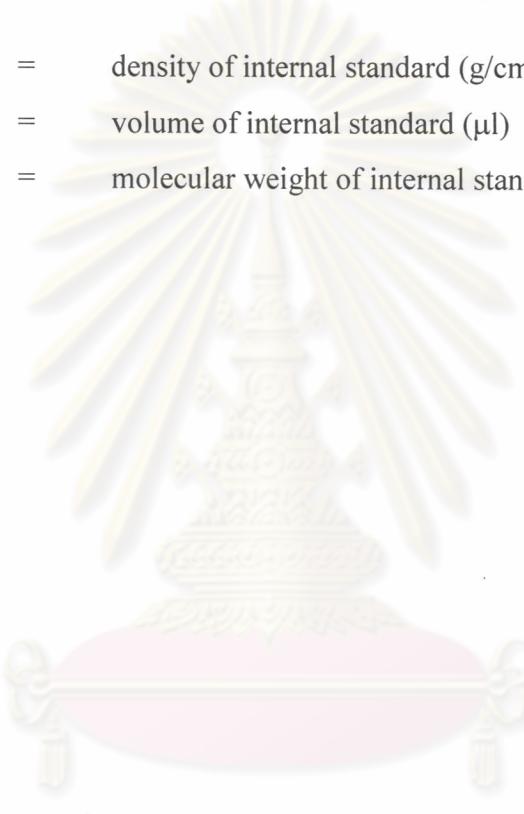
$A_{\text{interest}}$  = area of interested peak

$A_{\text{std}}$  = area of internal standard

$C_{\text{std}}$  = carbon effective number of the component in standard  
 $C_{\text{std}}$  = carbon effective number of the component in product  
 $\text{mol}_{\text{std}}$  = mol of internal standard

$$\text{mol}_{\text{std}} = \frac{d \times V \times 10^{-6} \times 1000}{\text{MW}}$$

$d$  = density of internal standard ( $\text{g/cm}^3$ )  
 $V$  = volume of internal standard ( $\mu\text{l}$ )  
 $\text{MW}$  = molecular weight of internal standard


  
**ศูนย์วิทยาศาสตร์พยาภรณ์**  
**จุฬาลงกรณ์มหาวิทยาลัย**

# VITAE

Mr. Tanawat Kanjanaboonmalert was born on October 28, 1979 in Bangkok, Thailand. He received a Bachelor Degree of Science in Chemistry from Chulalongkorn University in 2002. Since then, he has been a graduate student in major of Inorganic Chemistry, Faculty of Science, Chulalongkorn University. During his graduate studies towards the Master's Degree, he also receives a teaching assistantship from Department of Chemistry, Faculty of Science in 2003, and a research grant from the Graduate School, Chulalongkorn University. He also joins the international conferences and presents his thesis work in the poster form.

## Poster Topics

1. Kanjanaboonmalert, T.; Sukpirom, N.; Chaianansutcharit, S. "SYNTHESIS AND CATALYTIC ACTIVITY OF IRON OXIDES PILLARED HECTORITE", the 30<sup>th</sup> Congress on Science and Technology of Thailand, Bangkok, Thailand, 2004.
2. Kanjanaboonmalert, T.; Sukpirom, N.; Chaianansutcharit, S. "ALKYLATION OF BENZENE WITH 1-DODECENE OVER Fe-PILLARED CLAYS", the International Conference on Smart Materials-'04 (Smart/Intelligent Materials and Nanotechnology) Faculty of Science, Chiang Mai University, Chiang Mai, Thailand, 2004.