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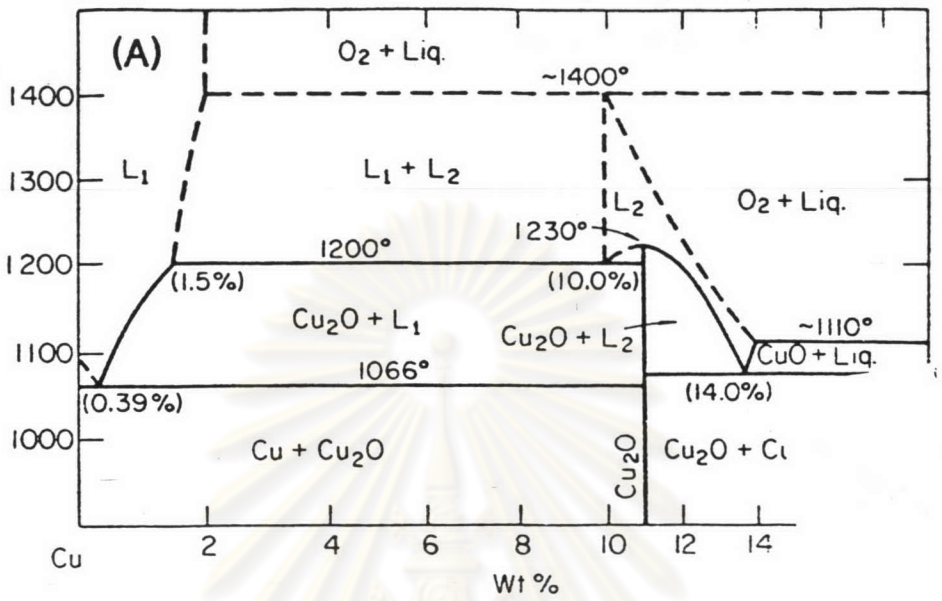
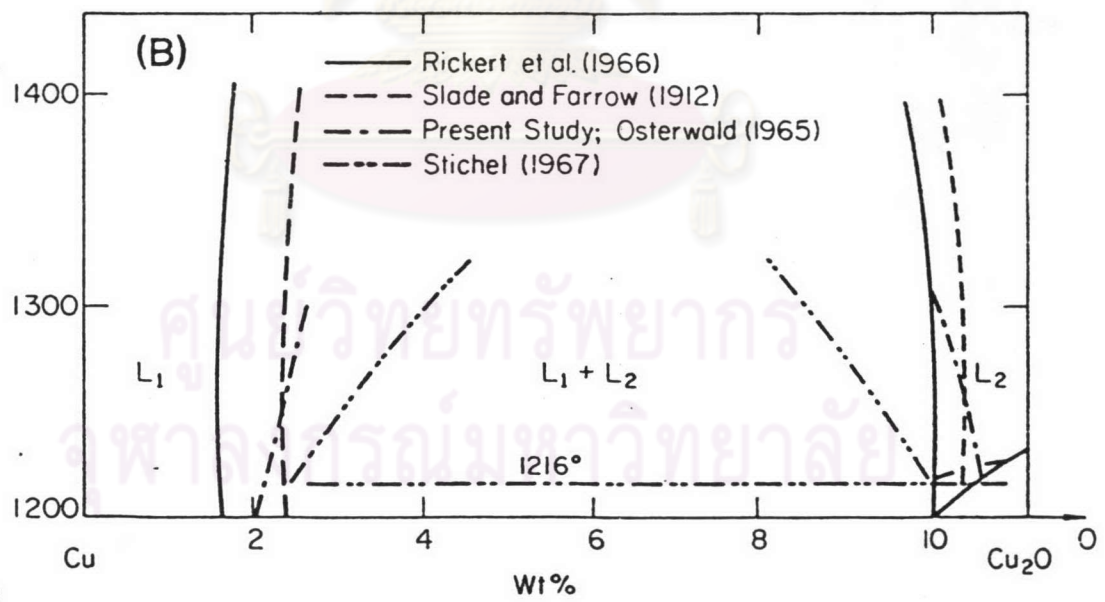
APPENDICES


(A)
O_2 + Liq.
~1400°
1400
1300
1200
1100
1000
L_1
$L_1 + L_2$
L_2
O_2 + Liq.
1230°
1200°
(1.5%)
(10.0%)
$Cu_2O + L_1$
$Cu_2O + L_2$
~1110°
CuO + Liq.
1066°
(0.39%)
(14.0%)
$Cu + Cu_2O$
Cu_2O
Cu_2O + Cu
Cu
2
4
6
8
10
12
14
Wt %


(B)
Rickert et al. (1966)
Slade and Farrow (1912)
Present Study; Osterwald (1965)
Stichel (1967)
1400
1300
1200
L_1
$L_1 + L_2$
L_2
1216°
Cu
2
4
6
8
10
Cu_2O
O
Wt %

Appendix B

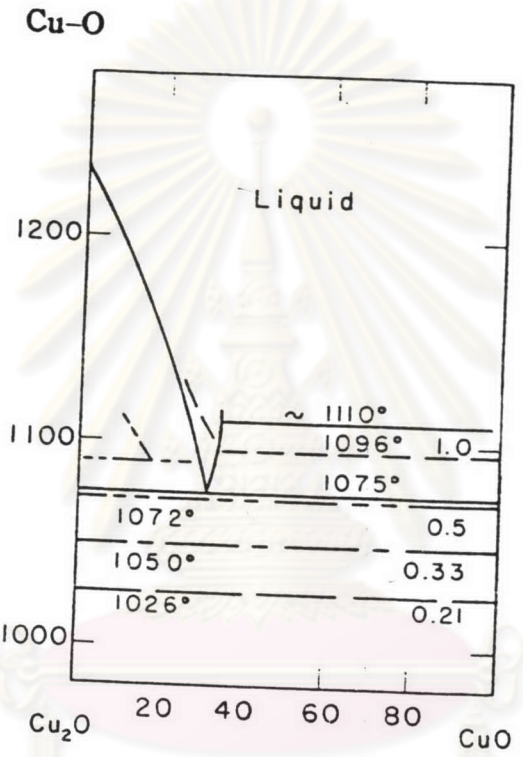


FIG. 2069.—System Cu_2O–CuO at various oxygen pressures. Solid lines are phase boundaries according to Vogel and Pöcher.

Appendix C

$G = a0 + a1 \cdot T + a2 \cdot T^2$ (*), T in 1000 K

	a0	a1	a2	r^2
Cu	-0.04	6.54	4.50	1.000
Cu2O	42.01	16.39	13.24	1.000
CuO	37.86	9.91	7.11	1.000
SiO2	215.90	10.58	8.17	1.000
SiO g	22.75	51.07	4.18	1.000
Si	0.26	2.48	4.52	0.999
O2	-0.72	49.54	4.02	1.000
H2	-0.71	31.84	3.64	1.000
H2O g	57.08	45.31	4.84	1.000

Constants a_o, a_1, a_2 for the calculation of the Gibbs free energies $G^\circ(T)$ in kcal/mol of selected species after equation (*) ; r^2 = squared regression coefficient; the constants are valid in the interval 298-2,000 K

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