



REFERENCES

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APPENDIX

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APPENDIX

Geochemical Techniques

X-ray Fluorescence method was mainly used for geochemical analysis of rocks for major element-oxides (Table 1) and trace elements (Table 2). The combined water (H_2O^+) and ferrous iron (FeO) were determined by a modified Penfield Method and volumetric method, respectively.

Preliminary Treatment of the Rock Sample

Fresh samples were carefully selected used for chemical analysis. The large sample was broken and pieces. Using a jaw crusher, the samples were crushed into small chips of about 2 - 3 mm. The crushed sample of about 100 gm was then pulverized to 200 - 250 mesh by a tungsten-carbide barrels on a crusher for 3 - 5 minutes. The powder sample was homogenized by mixing on the glazed paper and sampling by cone splitting.

Preparation of Glass Bead

1. Take 4 gm. of powder sample in an oven for two hours at 120 °C.
2. Weigh the dry sample.
3. Weigh the powder sample exactly 1.5 gm. in a platinum crucible (platinum/gold 95/5%). Add 7.5 gm. of a mixture of anhydrous spectromelt (mixing the equal amount of lithium metaborate

and lithium tetraborate flux to the sample with the ratio 5:1) and mix sample and spectromelt with a glass rod in the platinum crucible.

4. Put the platinum crucible on a Harzog burner for ten minutes at 1100 °C.

5. Take the bead out of the platinum crucible and write down the sample number after cooling has completed.

The glass bead was homogenized by the way of this method and used for the determination of major element oxides and trace elements by X-ray Fluorescence.

X-ray Fluorescence Method

The X-ray Fluorescence spectrometry is used for determining the element that present in a rock sample. The principle of this method is that when an element in a multi component sample is bombarded with primary X-ray photon, secondary X-ray photons characteristic of this element are produced. These secondary fluorescent X-rays are then dispersed according to their wave lengths. This is achieved by using a crystal of known d-spacing where diffraction takes place according to Bragg's relationship $n\lambda = 2d\sin\theta$ where λ is the wavelength, d is the spacing between the planes of the crystal. The intensity of the diffracted fluorescent radiation is then measured, and the concentration of the element producing this radiation can be found, since the intensity is proportional to concentration.

The sample analysis was carried out on a philips (PW. 1400) automatic X-ray spectrometer with computerized system at Analytical Laboratory Section, Geological Survey Division; Department of Mineral

Resources. The technique used in this work was the fusion method for glass-bead sample. The homogeneity and the surface of samples are to be in great case.

The regression analysis was applied also with the matrix corrections. These are carried out on onlined-phillips computer.

Determination of Combined Water (H_2O^+)

The combined water (H_2O^+) of the rock samples were determined by a modified Penfield Method. The procedure of the analytical technique is as follows:

1. Place 3 gm. of sodium tungstate flux and 1.000 gm. of the sample in a porcelain crucible and mix them thoroughly with a glass rod. Using a long-stemmed funnel, completely transfer this mixture into the bulb of a dried Penfield tube. Close the tube with a capillary stopper.

2. Horizontally insert the bulb of the Penfield tube, with a strip of damp cloth spirally wrapped around the middle sector, into the open flame from a bunzen burner. Slowly rotate the tube while heating the bulb at red-heated for 30 minutes. Keep the cloth wet throughout the heating period.

3. Heat the juncture of the tube until it softens. Pull off the bulb with iron tongs, and heat the sealed end of the tube for a few seconds to round of the sharp tip.

4. Let the tube cool, unwrap the damp cloth, and blot the water from the outside of the tube with a towel. Remove the stopper and weight the tube (tube weight plus weight of combined water).

5. Dry the tube in an oven at 100 °C for 1 hour. Let the tube cool and weight of the tube during drying.

Determination of Ferrous Iron

Ferrous iron was determined by volumetric method. The sample powder was treated with hot hydrofluoric and sulphuric acids. The contents of the platinum crucible were then transferred to a beaker containing a mixture of sulphuric, phosphoric and boric acids. The mixture was then titrated with potassium dichromate, until the indicator, sodium diphenylamine sulphonate assumed a permanent purple color. Fe_2O_3 was calculated from total Fe_2O_3 and FeO by using the relationship : $\text{Fe}_2\text{O}_3 = \text{total Fe}_2\text{O}_3 - (\text{FeO} \times 1.1113)$. This method was fast and direct, no comparisons were made with other rocks. The presence of sulphides or organic matter rendered an accurate determination of ferrous iron impossible. None of these components presented any serious problem in any of rocks analysed.

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BIOGRAPHY

Mr. Amarit Suvunsavate was born in Nakon Sri Thamarat, Thailand, on May, 25, 1947. He graduate with a B.Sc. in Geology from Chiangmai University, Chiangmai in 1970. Then he worked in Materials and Research Division, Department of Highways until 1975. During he worked in the said Division, he attempted to apply geological knowledge i.e. aerial photo interpretation, geophysical investigation to serve engineering work. From 1975 - 1984, he worked in Mining Division, Mines Organization. He served as Mining Geologist at Pilok mine, Chief of Exploration Section, Huai Yod Mine Manager, Changwat Trang and Senior Geologist, respectively. At present, he has worked as a Mining Geologist in the Viriya Industry Co. Ltd., Bangkok.

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