



REFERENCES

- Flego, C., Carluccio, L., Rizzo, C., and Perego, C. (2001). Synthesis of mesoporous $\text{SiO}_2\text{-ZrO}_2$ mixed oxides by sol-gel method. Catalysis Communications, 2, 43-48.
- Feng, Z., Postula, W., Akgerman, A., and Anthony, R. (1995). Characterization of Zirconia-Based Catalysts Prepared by Precipitation, Calcination, and Modified Sol-Gel Methods. Industry & Engineering Chemistry Research, 34, 78-82
- Gomez, R., Lopez, T., Tzompantzi, F., Garciafigueroa, E., Acosta, and Novaro, O (1997). Zirconia/Silica Sol-Gel Catalysts: Effect of the Surface Heterogeneity on the Selectivity of 2-Propanol Decomposition. Langmuir, 13, 970-973.
- Gonella, F., Mattei, G., and Mazzoldi, P (1999). Structural and optical Properties of Silver-Doped Zirconia and Mixed Zirconia-Silica Matrices Obtained by Sol-Gel Processing. Chem Mater, 11 (3), 814-821.
- Hino, M., and Arata, K. (1998). Catalysis of WO_3/ZrO_2 mechanically mixed with Pt/ZrO_2 for reaction of butane to isobutane. Applied Catalysis, 169, 151-155.
- Hardy, B., Walther, K. L., Wokuan, A., and Baiker, A. (1991). Preparation of catalysts. Amsterdam: Elsevier Science.
- Handy, B., Walther, K., Wokaun, A., and Baiker, A (1991). Influence of Preparation Parameters on Pore Structure of Silica Gels Prepared from Tetraethoxy Orthosilicate. Preparation of Catalyst V, 239-246.
- Hori, C.E., Brenner, A., Simon Ng, K.Y., Rahmoeller, K.M., and Belton, D. (1999). Studies of the oxygen release reaction in the platinum-ceria-zirconia system. Catalysis Today, 50, 299-308.
- Kuznetsov, P., Kuznetsova, L., Zhyzhaev, A., Pashkov, G., and Boldyrev, V (2002). Ultra fast synthesis of metastable tetragonal zirconia by means of mechanochemical activation. Applied Catalysis A:General, 227, 299-307.

- Ksapabutr, B., Wongkasemjit, S., and Gulari, E. A novel Synthesis of Zirconium and Cerium Glycolate Complexes. (Submitted)
- Laobutee, A., Wongkasemjit, S., Traversa, E., and Laine, R (2000). MgAl₂O₄ spinel powders from oxide one pot synthesis (OOPS) process for ceramic humidity sensors. Journal of the European Ceramic Society, 20, 91-97.
- Navio, J., Hidalgo, M., Colon, G., Botta, S., and Litter, M (2001). Preparation and Physicochemical Properties of ZrO₂ and Fe/ZrO₂ Prepared by a Sol-Gel Technique. Langmuir, 17, 202-210.
- Richardson, J. T., (1989). Principles of catalysts development. New York: Plenum Press.
- Rossignol, s., Madier, Y., and Duprez, D (1999) . Preparation of zirconia-ceria material by solf chemistry. Catalysis Today, 50, 261-270.
- Stocker, C., and Baiker, A. (1997). Zirconia aerogels: effect of acid-to-alkoxide ratio, alcoholic solvent and supercritical drying method on structural properties. J. of Non-crystalline Solids, 223, 165-178.
- Stocker, C., and Baiker, A. (1997). Zirconia aerogels: Effect of the Use of Mono- and Dicarboxylic Acids in the Sol-Gel Process on Structural Properties. J. of Sol-Gel Science and Technology, 10, 269-282.
- Sheng Wu, J., and Chuen Cheng, L. (2000). An improved synthesis of ultrafiltration zirconia membranes via the sol-gel route alkoxide precursor. J. of Membrane Science, 167, 253-261.
- Tanabe, K., Tan-No, M., and Yamaguchi, T (1991). ZrO₂ as a support: Oxidation of CO on CrO_x/ZrO₂. Preparation of Catalyst V, 567-574.
- Thammachart, M., Meeyoo, V., Risksomboon, T., and Osuwan, S. (2001). Catalytic activity of CeO₂-ZrO₂ mixed catalysts prepared via sol-gel technique: CO oxidation. Catalysis Today, 68, 63-61.
- Vacassy, R., Guizard, C., Thoraval, V., and Cot, (1997). Synthesis and characterization of microporous zirconia powders: Application in nanofilters and nanofiltration characteristics. Journal of Membrane Science, 132, 109-118.
- Ward, D., and Ko, E. (1993). Synthesis and Structural Transformation of Zirconia Aerogels. Chemistry of Materials, 5, 956-969.

- Ward, D., and Ko, E. (1994). Use of Preformed Sols in the Sol-Gel Preparation of Zirconia. Langmuir, 11, 369-372.
- Zhang, Q., Feng, Y., and Da, S. (1999). Prepration and Characterization of Zirconia-Silica and Zirconia-Magnesia Supports for Normal-Phase Liquid Chromatography. Analytical Science, 15, 767-772.

APPENDICES

Appendix A Raw Data

Table A1-1 Effect of pH on gel time obtained at the R_H of 587 and 1320

pH	Gel Time (s) at R_H 587	Gel Time (s) at R_H 1320
8	60	10
9	10	10
10	10	10
11	10800	9600
12	10800	9600
13	7200	3600
14	7200	3600

Table A1-2 Effect of the R_H on gel time obtained at pH 11-12 and 13-14

R_H	Gel time (s) at pH 11-12	Gel time (s) at pH 13-14
352	12000	9000
587	10800	8400
882	7200	4200
1320	9000	4200
1760	5400	4800

Table A1-3 Textural properties of zirconias calcined at 600° C

pH	R _H	S _{BET} (m ² /g)	V _p (N ₂) (cm ³ /g)	dp (Å)
7-8	352	24	0.056	86.7
	587	20	0.048	95.8
	882	19	0.047	
	1320	22	0.05	92
9-10	352	50	0.05	46
	587	55	0.057	40.3
	882	43	0.04	50.4
	1320	45	0.057	48.6
11-12	352	110	0.07	26
	587	125	0.077	24.8
	882	90	0.06	30.5
	1320	105	0.073	28.3
13-14	352	45	0.042	45.6
	587	50	0.055	43.7
	882	52	0.05	44.6
	1320	44	0.04	45.2

Note : R_H = molar water to precursor ratio

S_{BET} = BET surface area

V_p (N₂) = Total pore volume

dp = Average Pore diameter

Appendix B Calculations

B 1 Calculation of 100 % yield of catalyst

Using precursor = 1 gm

Mw. of Sodium Tris(glycozirconate) = 317

Mw. of zirconia = 123

Basis 1 gm of precursor

Zr mass balance

317 gm. of Sodium Tris(glycozirconate) contain 91 gm. of Zr

So 1 gm. of Sodium Tris(glycozirconate) contain 0.2870 gm. of Zr

100 % yield of catalyst

91 gm. of Zr can produce 123 gm. of ZrO_2

So 1 gm. of Sodium Tris(glycozirconate) that contain 0.2870 gm. of Zr
can be produce 0.3879 gm. of ZrO_2

100 % yield of ZrO_2 = 0.3879 gm.

Table B 1-1 Sodium removal from gel formed with the R_H of 587 at pH 9-10

Washing Times	Weight of cat.(gm)	% yield of cat.	%weight loss	% Na content
0	0.3563	100	0	10
3	0.3426	96.1549	3.8450	0.83
5	0.3220	90.3732	9.6267	0.68
7	0.3095	86.8650	13.1349	0.38
10	0.2921	81.9814	18.0185	0.35

Appendix C Experimental Conditions

C 1 Atomic Absorption Spectroscopy (AAS)

Preparation of standard solutions: 2.542 g of dried NaCl was dissolved in water and diluted to 1 litre to give 1000 ug/mL Na

Recommended of instrument parameters for atomic absorption experiments

Table C1-1 Working condition (FIXED)

Lamp current	5 mA
Fuel	acetylene
Support	air
Flame stoichiometry	oxidizing

Table C1-2 Working condition (VARIABLE)

Wavelength nm	Slit width nm	Optimum working range ug/mL
589	0.5	0.002-1.0
589.6	1.0	0.01-2.0
330.2 330.3	0.5	2-400

Table C1-3 Flame emission

Wavelength	589.0 nm
Slit width	0.1 nm
Fuel	acetylene
Support	air

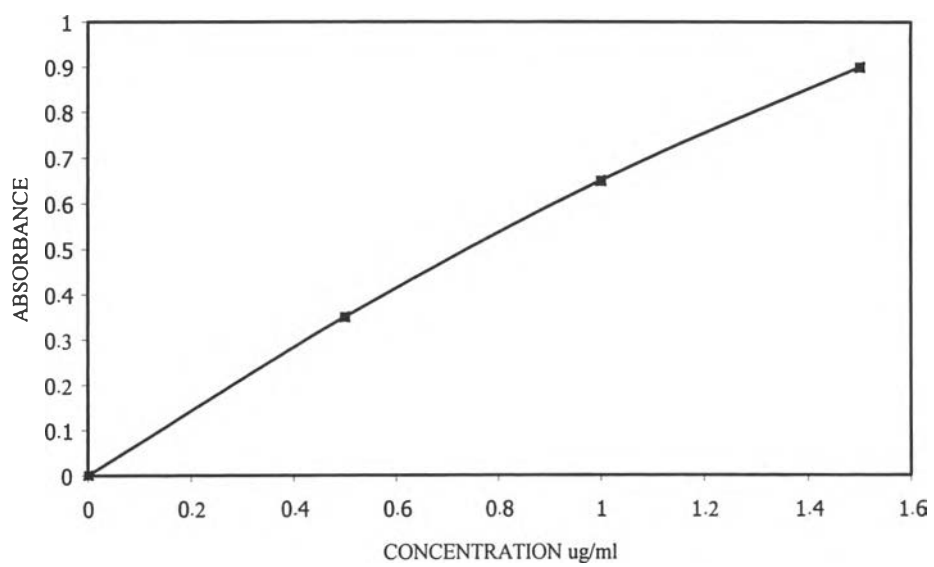


Figure C1-1 Calibration curve of sodium content

C 2 X-ray Diffraction (XRD)

JCPDS File name : Inorganic
 Card No. : 341084
 C. formular : ZrO_2
 C. or mineral name : Zr O
 Reli. Signs : I
 Target. : Cu Wave length : 1.54050 Ang

Table C 2-1 Standard peak of Zirconia from X-ray Diffraction (XRD)

Peak No.	d - value	Intensity	2 theta
1	2.925	100	30
2	2.623	75	34
3	2.509	50	36
4	1.791	100	50
6	1.704	25	54
7	1.641	25	56
8	1.565	50	59
9	1.518	75	61
10	1.462	50	64
11	1.409	25	66

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