

REFERENCES

- Adinata, D., Daud, W.M.A.W., and Aroua, M.K. (2007) Preparation and characterization of activated carbon from palm shell by chemical activation with K_2CO_3 . *Bioresource Technology*, 98, 145-149.
- Alcañiz-Monge, J., Lozano-Castelló, D., Cazorla-Amorós, D., and Linares-Solano, A. (2009) Fundamentals of methane adsorption in microporous carbons. *Microporous and Mesoporous Materials*, 124, 110-116.
- Arami-Niya, A., Daud, W.M.A.W., and Mjalli, F.S. (2011) Comparative study of the textural characteristics of oil palm shell activated carbon produced by chemical and physical activation for methane adsorption. *Chemical Engineering Research and Design*, 89, 657-664.
- Azevedo, D.C.S., Araújo, J.C.S., Bastos-Neto, M., Torres, A.E.B., Jaguaribe, E.F., and Cavalcante, C.L. (2007) Microporous activated carbon prepared from coconut shells using chemical activation with zinc chloride. *Microporous and Mesoporous Materials*, 100, 361-364.
- Balathanigaimani, M.S., Kang, H-C., Shim, W-G., Kim, C., Lee, J-W., and Moon, H. (2006) Preparation of powdered activated carbon from rice husk and its methane adsorption properties. *Korean Chemical Engineering*, 23 (4), 663-668.
- Brunauer, S., Deming, L., Deming, W., and Teller, E.J. (1940) On a theory of the van der Waals Adsorption of gases. *Journal of the American Chemical Society*, 1940, 62 (7), 1723-1732.
- Burchell, T. and Rogner, M. (2000) Low Pressure storage of natural gas for vehicular applications. *Society of Automotive Engineers, Inc*, 01-2205.
- Daud, W.M.A.W., and Ali, W.S.W. (2004) Comparison on pore development of activated carbon produced from palm shell and coconut shell. *Bioresource Technology*, 93, 63-69.
- Daud, W.M.A.W., Ali, W.S.W., and Sulaiman, M.Z. (2000) The effects of carbonization temperature on pore development in palm-shell- and based activated carbon. *Carbon*, 38, 1925-1932.

- Delavar, M., Ghoreyshi, A.A., Jahanshahi, M., and Irannejad, M. (2010) Experimental evaluation of methane adsorption on granular activated Carbon (GAC) and Determination of model isotherm. Engineering and Technology, 62, 47-50.
- Dreisbach, F., Staudt, R., and Keller, J.U. (1999) High Pressure adsorption data of methane, nitrogen, carbon dioxide and their binary and ternary mixtures on activated carbon. Adsorption, 5, 215-227.
- Giraldo, L., and Moreno-Piraján, J.C. (2011) Novel activated carbon monoliths for methane adsorption obtained from coffee husks. Materials Sciences and Applications, 2, 331-339.
- Hayashi, J., Horikawa, T., Takeda, I., Muroyana, K., and Ani, F.N. (2002) Preparing activated carbon from various nutshells by chemical activation with K_2CO_3 . Carbon, 40, 2381-2386.
- Hernández-Montoya, V., García-Servin, J., and Bueno-López, J.I.. (2012) Lignocellulosic Precursors Used in the Synthesis of Activated Carbon-Characterization Techniques and Applications in the Wastewater Treatment. InTech. (pp 19-36).
- IUPAC (International Union of Pure and Applied Chemistry) Recommendation. Reporting physisorption data for gas/solid systems with special reference to the determination of surface area and porosity, Pure and Applied Chemistry, 1985, 57 (4), pp.603-619.
- Jia, Q., and Lua, A.C. (2008) Effects of pyrolysis conditions on the physical characteristics of oil-palm-shell activated carbons used in aqueous phase phenol adsorption. Journal of Analytical and Applied Pyrolysis, 83, 175-179.
- Jia, Z., Li, H., Yu, Z., Wang, P., and Fan, X. (2011) Densification of MOF-5 synthesized at ambient temperature for methane adsorption. Materials Letters, 65, 2445-2447.
- Kumpoomee, N. (2012) Comparative study for methane adsorption on metal organic frameworks and activated carbon, M.S. Thesis, The Petroleum and Petrochemical College, Chulalongkorn University, Bangkok, Thailand.

- Li, J., Cheng, S., Zhao, Q., Long, P., and Dong, J. (2009) Synthesis and hydrogen-storage behavior of metal-organic framework MOF-5. International Journal of Hydrogen Energy, 34, 1377-1382.
- Li, W., Yang, K., Peng, J., Zhang, L., Guo, S., and Xia, H. (2008) Effects of carbonization on temperatures on characteristics of porosity in coconut shell chars and activated carbons derived from carbonized coconut shell chars. Industrial Crops and Products, 28, 190-198.
- Lozano-Castelló, D., Cazorla-Amorós, D., and Linares-Solano, A. (2002) Powdered activated carbons and activated carbon fibers for methane storage: a comparative study. Energy & Fuels, 16, 1321-1328.
- Lozano-Castelló, D., Cazorla-Amorós, D., Linares-Solano, A., and Quinn, D.F. (2002) Influence of pore size distribution on methane storage at relatively low pressure: preparation of activated carbon with optimum pore size. Carbon, 40, 989-1002.
- Lozano-Castelló, D., Lillo-Ródenas, M.A., Cazorla-Amorós, D., and Linares-Solano, A. (2001) Preparation of activated carbons from Spanish anthracite: I. activation by KOH. Carbon, 39, 741-749.
- Matranga, K.R., Myers, A.L., and Glandt, E.D. (1992) Storage of natural gas by adsorption on activated carbon. Chemical Engineering Science, 47 (7), 1569–1579.
- Méndez-Liñán, L., López-Garzón, F.J., Domingo-García, M. and Pérez-Mendoza, M. (2010) Carbon adsorbents from Polycarbonate pyrolysis char residue: hydrogen and methane storage capacities. Energy & Fuels, 24 (6), 3394–3400.
- Monocha, S.M. (2003) Porous carbons. Sādhanā, 28, 335-348.
- Mi, J., Wang, X.R., Fan, R.J., Qu, W.H., and Li, W.C. (2012) Coconut-shell-based porous carbons with a tunable micro/mesopore ratio for high-performance supercapacitors. Energy & Fuels, 26 (8), 5321-5329.
- Najibi, H., Chapoy, A., and Bahman, T. (2008) Methane/natural gas storage and delivered capacity for activated carbons in dry and wet conditions. Fuel, 87, 7-13.

- Qada, E.N.E., Allen, S.J., and Walker, G.M. (2006) Adsorption of methylene blue onto activated carbon produced from steam activated bituminous coal: a study of equilibrium adsorption isotherm. *Chemical Engineering Journal*, 124, 103-110.
- Rouquerol, F., Rouquerol, J., and Sing, K. (1999) *Adsorption by Powder & Porous Solid*. London: Academic Press.
- Salehi, E., Taghikhani, V., Ghotbi, C., Lay, E.N., and Shojaei, A. (2007) Theoretical and experimental study on the adsorption and desorption of methane by granular activated carbon at 25°C. *Journal of Natural Gas Chemistry*, 16, 415–422.
- Sing, K.S.W. (1982) Reporting physisorption data for gas/solid systems with special reference to the determination of surface area and porosity. *Pure and Applied Chemistry*, 54 (11), 2201-2218.
- Solar, C., Blanco, A.G., Vallone, A., and Sapag, K. (2010) Natural gas. *Journal of Natural Gas Chemistry*, pp 205–244, InTech.
- Sun, Y., Yang, G., Wang, Y.S., and Zhang, J.P. (2011) Production of activated carbon by K₂CO₃ activation treatment of furfural production waste and its application in gas storage. *Environment Progress & Sustainable Energy*, 30 (4), 648–657.
- Tam, M.S. and Antal, M.J., Jr. (1999) Preparation of activated carbons from macadamia nut shell and coconut shell by air activation. *Industrial & Engineering Chemistry Research*, 38, 4268–4276.
- Tan, I.A.W., Ahmad, A.L., and Hameed, B.H. (2008) Adsorption of basic dye on high-surface-area activated carbon prepared from coconut husk: equilibrium, kinetic and thermodynamic studies. *Journal of Hazardous Materials*, 154, 337–346.
- Tranchemontagne, D.J., Hunt, J.R., and Yaghi, O.M. (2008) Room temperature synthesis of metal-organic frameworks: MOF-5, MOF-74, MOF-177, MOF-199, and IRMOF-0. *Tetrahedron*, 64, 8553-8557.

- Tsai, W.T., Chang, C.Y., Wang, S.Y., Chang, C.F., Chien, S.F., and Sun, H.F. (2001) Preparation of activated carbons from corn cob catalyzed by potassium salts and subsequent gasification with CO₂. Bioresource Technology, 78, 203–208.
- Wutthikun, P. (2001) Effect of activation by carbon dioxide and superheated steam on properties of activated carbon from coconut shells, M.S. Thesis, Faculty of Science, Chulalongkorn University, Bangkok, Thailand.
- Yusufu, M.I., Ariahu, C.C., and Igbabul, B.D. (2012) Production and characterization of activated carbon from selected local raw materials. African Journal of Pure and Applied Chemistry, 6 (9), 123-131.
- “Adsorption” 25 March 2013
<<http://www.chemvironcarbon.com>>
- “Liquefied natural gas” 5 June 2012
<http://www.pttln.com/en/mr_lng.aspx>
- “Natural Gas” 8 June 2012
<<http://www.naturalgas.org>>

APPENDICES

Appendix A Specifications of the adsorbents

Table A1 Specifications of the commercial activated carbons

Activated carbon	Iodine Number (mg/g)	Moisture (%)	Ash (%)	Hardness Number (%)	Apparent Density (g/cc)
A (IN1100)	1100	8.00	3.50	98.00	0.48
B (IN 1067)	1067	1.30	6.00	97.50	0.51
C (IN 1100)	1100	8.00	5.00	98.00	0.48
D (IN 937)	973	2.70	7.50	96.00	0.49
E (IN 1035)	1035	1.70	7.30	98.00	0.48
F (IN 879)	879	1.44	8.90	97.00	0.49
G (IN 987)	987	1.70	8.70	96.10	0.51

Table A2 Physical properties of adsorbents (Kumpoomee, 2012)

Adsorbents	BET surface area (m ² /g)	V _{DR} (cc/g)	Total pore volume (cc/g)	DR micropore (Å)	Average pore diameter (Å)
Basolite C300	2434	1.27	1.33	13.9	21.9
Basolite Z1200	1449	0.78	0.88	12.1	24.3
Calgon	877.8	0.48	0.53	22.3	24.3
Coconut Shell Granular Activated Carbon	865.4	0.47	0.49	18.9	22.3
Eucalyptus Powder Activated Carbon	861	0.47	0.61	25.2	28.2
Coconut Shell Powder Activated Carbon	713.5	0.39	0.42	17.1	23.5

Appendix B The Amount of Methane Adsorbed on all Adsorbents

Table B1 The amount of methane adsorption on A (IN 1100) at 35 °C

Equilibrium pressure (psia)	Methane adsorption (mmol/g)
15.00	0
48.75	2.317
99.38	3.337
150.00	3.981
202.50	4.379
300.00	4.977
401.25	5.293
498.75	5.583
600.00	5.746
696.38	5.922
801.88	6.121
900.00	6.230
1000.00	6.282

Table B2 The amount of methane adsorption on A (IN 1100) at 40 °C

Equilibrium pressure (psia)	Methane adsorption (mmol/g)
13.13	0
50.53	2.251
101.26	3.216
150.01	3.865
202.51	4.285
300.01	4.854
401.26	5.189
500.63	5.381
598.13	5.594
695.63	5.713
800.63	5.855
898.26	5.930
998.51	5.992

Table B3 The amount of methane adsorption on A (IN 1100) at 45 °C

Equilibrium pressure (psia)	Methane adsorption (mmol/g)
15.00	0
54.38	2.063
103.13	3.002
153.75	3.524
202.50	3.974
301.88	4.451
401.25	4.772
496.88	5.106
598.38	5.278
693.75	5.452
802.50	5.534
900.00	5.636
997.50	5.682

Table B4 The amount of methane adsorption on B (IN 1067) at 35 °C

Equilibrium pressure (psia)	Methane adsorption (mmol/g)
15.00	0
50.63	1.773
100.63	2.611
148.13	3.177
199.38	3.538
294.38	4.015
399.38	4.413
500.00	4.649
600.00	4.836
697.50	4.966
799.38	5.066
898.38	5.166
1000.00	5.212

Table B5 The amount of methane adsorption on B (IN 1067) at 40 °C

Equilibrium pressure (psia)	Methane adsorption (mmol/g)
15.00	0
50.63	1.771
101.25	2.585
150.00	3.072
200.00	3.475
301.88	4.012
403.13	4.270
502.50	4.453
600.00	4.618
701.25	4.785
802.50	4.909
896.25	4.979
995.63	5.157

Table B6 The amount of methane adsorption on B (IN 1067) at 45 °C

Equilibrium pressure (psia)	Methane adsorption (mmol/g)
15.00	0
54.38	1.643
105.00	2.372
151.88	2.869
200.63	3.337
301.88	3.902
403.75	4.189
503.25	4.393
602.50	4.534
703.13	4.668
804.38	4.765
896.25	4.863
997.50	4.956

Table B7 The amount of methane adsorption on C (IN 1100) at 35 °C

Equilibrium pressure (psia)	Methane adsorption (mmol/g)
15.00	0
48.75	2.116
99.38	2.999
150.00	3.485
198.75	3.894
298.13	4.343
397.50	4.630
498.75	4.834
598.13	4.959
695.63	5.067
800.63	5.094
898.25	5.142
999.38	5.161

Table B8 The amount of methane adsorption on C (IN 1100) at 40 °C

Equilibrium pressure (psia)	Methane adsorption (mmol/g)
16.88	0
52.51	2.035
103.13	2.850
150.01	3.371
200.63	3.705
300.01	4.193
399.38	4.497
500.63	4.708
600.01	4.826
697.51	4.872
802.51	4.885
899.13	4.899
999.38	4.918

Table B9 The amount of methane adsorption on C (IN 1100) at 45 °C

Equilibrium pressure (psia)	Methane adsorption (mmol/g)
16.88	0
54.38	2.001
103.13	2.790
153.76	3.248
206.26	3.564
304.38	4.077
405.01	4.366
504.38	4.633
607.51	4.733
704.13	4.787
803.76	4.866
901.26	4.855
1003.26	4.868

Table B10 The amount of methane adsorption on D (IN 937) at 35 °C

Equilibrium pressure (psia)	Methane adsorption (mmol/g)
15.00	0
48.75	1.630
99.38	2.274
148.13	2.752
196.88	3.132
296.25	3.570
395.63	3.940
495.00	4.163
594.38	4.396
697.50	4.532
796.88	4.678
892.50	4.837
995.63	4.938

Table B11 The amount of methane adsorption on D (IN 937) at 40 °C

Equilibrium pressure (psia)	Methane adsorption (mmol/g)
16.88	0
52.51	1.549
103.13	2.312
155.63	2.745
204.38	3.083
298.13	3.464
403.13	3.858
504.38	4.072
600.01	4.320
701.88	4.490
800.03	4.665
901.88	4.757
1001.88	4.881

Table B12 The amount of methane adsorption on D (IN 937) at 45 °C

Equilibrium pressure (psia)	Methane adsorption (mmol/g)
15.00	0
54.38	1.566
103.13	2.207
150.00	2.721
202.25	3.040
301.88	3.429
406.13	3.769
501.88	3.999
601.38	4.182
703.75	4.372
803.25	4.499
901.38	4.601
1001.25	4.725

Table B13 The amount of methane adsorption on E (IN 1035) at 35 °C

Equilibrium pressure (psia)	Methane adsorption (mmol/g)
15.00	0
50.63	1.633
97.50	2.410
148.13	2.924
198.75	3.283
296.25	3.787
399.38	4.099
500.00	4.315
600.88	4.496
698.75	4.665
798.25	4.748
900.00	4.783
1000.38	4.841

Table B14 The amount of methane adsorption on E (IN 1035) at 40 °C

Equilibrium pressure (psia)	Methane adsorption (mmol/g)
15.00	0
50.63	1.468
99.38	2.167
148.25	2.645
200.63	2.966
296.25	3.456
399.38	3.907
498.75	4.124
598.13	4.358
695.63	4.552
800.13	4.664
896.25	4.736
997.50	4.758

Table B15 The amount of methane adsorption on E (IN 1035) at 45 °C

Equilibrium pressure (psia)	Methane adsorption (mmol/g)
15.00	0
50.63	1.487
101.25	2.166
151.88	2.577
202.50	2.885
299.38	3.315
402.25	3.682
499.88	3.963
600.00	4.200
697.50	4.405
796.38	4.540
897.50	4.633
1000.00	4.689

Table B16 The amount of methane adsorption on F (IN 879) at 35 °C

Equilibrium pressure (psia)	Methane adsorption (mmol/g)
16.88	0
48.76	1.647
99.38	2.361
151.88	2.839
200.63	3.183
298.13	3.651
400.01	4.015
499.38	4.246
596.88	4.411
700.26	4.499
798.76	4.600
896.26	4.666
995.63	4.816

Table B17 The amount of methane adsorption on F (IN 879) at 40 °C

Equilibrium pressure (psia)	Methane adsorption (mmol/g)
15.00	0
48.75	1.610
97.50	2.340
150.00	2.808
200.63	3.153
296.25	3.626
398.13	4.000
500.63	4.196
598.13	4.382
701.25	4.430
800.63	4.499
898.13	4.585
997.50	4.710

Table B18 The amount of methane adsorption on F (IN 879) at 45 °C

Equilibrium pressure (psia)	Methane adsorption (mmol/g)
15.00	0
52.50	1.457
101.25	2.163
151.88	2.626
200.63	2.993
298.13	3.427
398.38	3.806
499.38	3.978
595.00	4.199
697.50	4.281
797.25	4.329
897.50	4.337
997.50	4.390

Table B19 The amount of methane adsorption on G (IN 987) at 35 °C

Equilibrium pressure (psia)	Methane adsorption (mmol/g)
16.88	0
52.51	1.527
99.38	2.274
150.01	2.749
202.51	3.147
300.01	3.688
403.26	3.972
502.51	4.226
602.51	4.399
704.38	4.495
802.63	4.606
901.26	4.646
1003.13	4.743

Table B20 The amount of methane adsorption on G (IN 987) at 40 °C

Equilibrium pressure (psia)	Methane adsorption (mmol/g)
15.00	0
52.50	1.520
101.25	2.185
150.00	2.636
202.50	2.960
298.75	3.384
399.38	3.712
498.75	3.939
601.88	4.047
701.25	4.167
800.00	4.278
896.25	4.385
998.13	4.488

Table B21 The amount of methane adsorption on G (IN 987) at 45 °C

Equilibrium pressure (psia)	Methane adsorption (mmol/g)
15.00	0
52.50	1.435
99.38	2.072
148.13	2.518
200.63	2.852
298.13	3.324
399.38	3.698
500.63	3.863
600.00	4.034
701.13	4.113
798.13	4.253
896.13	4.300
998.38	4.342

Table B22 The amount of methane adsorption on Co-K₂CO₃/1h at 35 °C

Equilibrium pressure (psia)	Methane adsorption (mmol/g)
15.00	0
33.75	1.214
90.00	1.679
148.13	1.908
200.63	2.034
301.88	2.148
399.38	2.249
498.75	2.287
599.38	2.326
697.50	2.384
798.25	2.411
899.38	2.430
1001.25	2.443

Table B23 The amount of methane adsorption on Co-K₂CO₃/1h at 40 °C

Equilibrium pressure (psia)	Methane adsorption (mmol/g)
15.00	0
46.88	0.740
95.63	1.164
150.00	1.401
198.75	1.612
300.00	1.787
399.38	1.899
498.75	1.962
600.00	2.007
699.38	2.031
798.75	2.046
900.00	2.055
1001.25	2.078

Table B24 The amount of methane adsorption on Co-K₂CO₃/1h at 45 °C

Equilibrium pressure (psia)	Methane adsorption (mmol/g)
15.00	0
48.75	0.668
97.50	1.046
150.00	1.302
200.63	1.457
301.88	1.605
401.25	1.693
498.75	1.796
600.00	1.870
700.00	1.901
800.00	1.921
900.75	1.935
1003.13	1.972

Table B25 The amount of methane adsorption on Co-K₂CO₃/2h at 35 °C

Equilibrium pressure (psia)	Methane adsorption (mmol/g)
15.00	0
33.75	1.183
87.25	1.741
146.33	2.002
200.63	2.109
301.88	2.246
399.38	2.369
500.63	2.404
600.00	2.456
699.38	2.477
798.75	2.502
900.00	2.520
1001.25	2.552

Table B26 The amount of methane adsorption on Co-K₂CO₃/2h at 40 °C

Equilibrium pressure (psia)	Methane adsorption (mmol/g)
15.00	0
41.25	0.948
88.13	1.577
148.13	1.796
200.63	1.930
301.88	2.076
399.38	2.220
499.38	2.294
600.00	2.364
699.38	2.413
799.38	2.457
900.63	2.474
1001.25	2.490

Table B27 The amount of methane adsorption on Co-K₂CO₃/2h at 45 °C

Equilibrium pressure (psia)	Methane adsorption (mmol/g)
15.00	0
48.75	0.730
97.50	1.151
151.88	1.376
202.50	1.509
301.88	1.719
401.25	1.842
498.75	1.983
601.88	2.022
701.25	2.074
800.63	2.132
900.00	2.173
1005.00	2.197

Table B28 The amount of methane adsorption on Co-K₂CO₃/3h at 35 °C

Equilibrium pressure (psia)	Methane adsorption (mmol/g)
15.00	0
41.25	0.931
93.75	1.345
150.00	1.542
200.63	1.677
301.25	1.802
399.38	1.908
498.75	1.977
600.00	2.024
699.38	2.057
798.75	2.100
900.00	2.127
1001.25	2.163

Table B29 The amount of methane adsorption on Co-K₂CO₃/3h at 40 °C

Equilibrium pressure (psia)	Methane adsorption (mmol/g)
15.00	0
50.63	0.590
99.38	0.924
151.88	1.142
200.63	1.311
301.88	1.460
401.25	1.564
498.75	1.674
601.88	1.700
701.25	1.736
800.63	1.777
900.00	1.874
1003.13	1.916

Table B30 The amount of methane adsorption on Co-K₂CO₃/3h at 45 °C

Equilibrium pressure (psia)	Methane adsorption (mmol/g)
15.00	0
61.88	0.2482
105.00	0.5366
153.75	0.7608
204.38	0.8489
301.88	1.0603
401.25	1.1834
500.63	1.2636
601.88	1.3196
701.25	1.3481
800.00	1.4054
901.88	1.4248
1005.00	1.4447

CURRICULUM VITAE

Name: Ms.Wasa Sunthonsuriyawong

Date of Birth: March 14, 1988

Nationality: Thai

University Education:

2007–2010 Bachelor Degree of Science, Chemical Technology (Fuel Technology), Chulalongkorn University, Thailand

Work Experience:

2010 Position: Internship

Company: PTT Aromatics and Refining Public Company
Limited, Rayong, Thailand

Proceedings:

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