

REFERENCES

- Amiri, M.C. and Valsaraj, K.T. (2004) Effect of gas transfer on separation of whey protein with aphon flotation. Separation Purification Technology, 35, 161-167.
- Bakarat, Y., Fortney, L.N., Lalanne, C.C., Schechter, R.S., Wade, W.H., Weerasooriya, U. and Yiv, S. (1983) The Phase Behavior of Sample Salt-Tolerant Sulfonates, Social Petroleum Engineering Journal, 23(1), 913-918.
- Binks, B.P., Cho, W-G., Fletcher and Petsev, D.N. (2000) Stability of Oil-in-Water Emulsions in a Low Interfacial Tension System, Langmiur, 16(3), 1025-1034
- Bourrel, M. and Schechter, R.S. (1988) Microemulsions and Related Systems. New York: Marcel Dekker.
- Carleson, T.E. (1989) Adsorptive Bubble Separation Processes. In Scamehorn, J.F., and Harwell, J.H. (Eds.), Surfactant Based Separation Process. New York: Marcel Dekker.
- Chaphalkar, P. G., Valsaraj, K. T. and Roy, D. (1993) A study of the size distribution and stability of Colloidal Gas Aphrons using a particle size analyser. Separation Science and Technology, 28, 1287-1302.
- Cilliers, J.J. and Bradshaw, D.J.(1996) The flotation of fine pyrite using colloidal gas aphonrs. Minerals Engineering, 9(2), 235-241
- Clarence, A.M. and Neogi, P. (1985) Interfacial Phenomena: Equilibrium and Dynamic Effects. New York: Marcel Dekker.
- Feng, D. and Aldrich, C. (2000) Removal of diesel from aqueous emulsions by flotation. Separation Science and Technology, 35, 2159-2172.
- Gu, X. and Chiang, S.(1999) A novel flotation column for oily wastewater clean up. Separation Purification Technology, 16,193-203.
- Hashim, M.A., Day, A., Hasan, S. and Sen Gupta, B. (1999) Mass transfer correlation in flotation of plam oil by colloidal gas aphonrs. Bioprocess Engineering, 21, 401-404.

- Hashim, M.A. and Sen Gupta, B. (1998) The application of colloidal gas aphrons in the recovery of fine cellulose fibers from paper mill wastewater. Bioresource Technology, 64, 199-204.
- Jarudilokkul, S., Rungphetcharat, K. and Boonamnuayvitaya, V. (2004) Protein separation by colloidal gas aphrons using nonionic surfactant. Separation and Purification Technology, 35, 23-29.
- Jauregi, P., Gilmour, S. and Varley, J. (1997) Characterisation of colloidal gas aphrons for subsequent use for protein recovery. Chemical Engineering Journal, 65, 1-11.
- Moosai, R. and Dawe, R.A. (1989) Oily wastewater cleanup by gas flotation. West Indian Journal Engineering, 25(1), 25-41.
- Pal, R. and Maliyah, J. (1990) Oil Recovery from Oil in Water Emulsions Using a flotation column, Canadian Journal Chemical Engineering, 68, 959-967.
- Patterson, J.W. (1975) Wastewater Treatment Technology. Michican: Ann Arbor.
- Phoochinda, W. (1997) Removal of emulsified oil from wastewater using froth flotation. M S Thesis, The Petroleum and Petrochemical College, Chulalongkorn University, Bangkok, Thailand.
- Pongstabodee, S., Scamehorn, J.F., Chavadej, S., Saiwan, C. and Harwell, J.H. (1998). Removal ortho-dichlorobenzene by froth flotation under Winsor's type III conditions. Separation Science and Technology, 23(4), 591-609.
- Puerto, M.C. and Reed, R.L. (1983) A Three-Parameter Representation of Surfactant/Oil/Brine Interaction, Social Petroleum Engineering Journal, 23(1), 669-682.
- Ramirez, J.A. and Davis, R.H. (2001) Microflotation of fine oil droplets by small air bubbles: experiment and theory. Separation Purification Technology, 36(1), 1-15.
- Ratanarojanatam, P. (1995) Clean-up of oily wastewater by froth flotation: effect of microemulsion formation by surfactant mixture. M.S. Thesis, The Petroleum and Petrochemical College, Chulalongkorn University, Bangkok, Thailand.

- Rosen, M.J., (1989) Surfactant and Interfacial Phenomena New York: John Wiley.
- Roy, D., Valsaraj, K.T. and Kottai, S.A. (1992) Separation of organic dyes from wastewater by using colloidal gas aphrons. Separation Science and Technology, 25(5), 573-588.
- Rubio, J., Souza, M.L., Smith, R.W. (2002) Overview of flotation as a wastewater treatment technique. Minerals Engineering, 15, 139-155.
- Save, S.V., Pangarkar, V.G. and Kumar, S.V. (1993) Intensification of mass transfer in aqueous two-phase systems. Biotechnology and Bioengineering, 41, 72-78.
- Save, S.V. and Pangarkar, V.G. (1994) Characterization of CGAs. Chemical Engineering Communication, 127, 35-54.
- Scamehorn, J.F. and Harwell, J.H. (2000) Surfactant Based Separation Process. New York: Marcel Dekker.
- Sebba, F. (1987) Colloidal gas aphrons, Foam and bilyquid foam-aphrons. Chichester: Wiley.
- Sebba, F. (1989) Novel Separations Using Aphrons. In Scamehorn, J.F., and Harwell, J.H. (Eds.), Surfactant Based Separation Process. New York: Marcel Dekker.
- Sharma, M.K. (1991) Particle Technology and Surface Phenomena, 2nd Ed. New York: Wiley.
- Sylvester, N.D. and Byeseda, J.J. (1980) Oil /Water Separation by Induced Air Flotation. Social Petroleum Engineering Journal, 20(6), 579
- Phoochinda, W. (1997) Removal of emulsified oil from wastewater using froth flotation. M.S Thesis, The Petroleum and Petrochemical College, Chulalongkorn University, Bangkok, Thailand.
- Watcharasing, S. (2004) Diesel Removal by Continuous Froth Flotation : Effects of Ultralow Interfacial Tension and Foam Characteristics. M.S Thesis, The Petroleum and Petrochemical College, Chulalongkorn University, Bangkok, Thailand.

- Withayapanyanon, A. (2003) Microemulsion Formation of Surfactant/Oily Wastewater System Related to Clean-Up by Froth Flotation. M.S. Thesis, The Petroleum and Petrochemical College, Chulalongkorn University, Bangkok, Thailand.
- Wu, B., Harwell, J.H., Sabatini, D.A. and Bailey, J.D. (2000) Alcohol-Free DIPHENYL Oxide Disulfonate Middle-Phase Microemulsion Systems, Journal of Surfactants and Detergents, 3(4), 465-474.
- Wungrattanasopon, S., Scamehorn, J.F., Chavadej. S., Saiwan, C. and Harwell, J.H. (1996) Use of foam flotation to remove *tert*-butylphenol from water. Separation Science and Technology, 31(11), 1523-1540.

APPENDICES

Appendix A Experimental Data of Microemulsion Formation.

1. Interfacial Tension (IFT)

The interfacial tension of each phase of microemulsion is interpreted by the following formulation:

$$\text{IFT} = e(Vd)^3 n^2 \Delta\rho \quad (\text{A1})$$

where

σ = interfacial tension or IFT (mN/m, dyne/cm)

e = unity factor ($3.427 \cdot 10^{-7}$ mN cm³ min² /m g mm³)

V = enlargement factor (0.31 mm/sdv)

d = measured drop diameter (sdv)

n = number of revolution (1/min)

$\Delta\rho$ = density difference of two liquids (g/cm³)

2. Experimental Data of Interfacial Tension (IFT)

Table A1 Interfacial tension of each phase in microemulsion formation with 0.05 wt% of Alfoterra at different NaCl concentrations and an initial oil to water ratio = 1:1

Alfoterra conc. (wt%)	NaCl conc. (wt%)	No.	Upper density (g/mL)	Lower density (g/mL)	Upper level	Lower level	Speed (rpm)	IFT (mN/m)
0.05	2	1	0.8160	0.9880	4.5900	2.6700	3834	3.87970
		2	0.8130	0.9870	4.6400	2.5500	3744	4.77201
		3			4.6900	2.5700	3665	4.77250
		ave	0.8145	0.9875	4.6400	2.5967		4.47474
	3	1	0.8/13	0.9990	4.5200	2.5700	3910	4.07400
		2	0.8090	0.9940	4.4700	2.6000	3878	3.53430
		3			4.5600	2.5800	3500	3.41739
		ave	0.8090	0.9965	4.5167	2.5833		3.80415
	4	1	0.8080	1.0060	4.5400	2.9900	3708	1.85098
		2	0.8040	1.0020	4.5500	2.8000	3566	2.46379
		3			4.4300	2.8500	3939	2.21243
		ave	0.8060	1.0040	4.5450	2.8800		2.17573
	5	1	0.8090	1.0020	4.7100	2.6100	2855	2.87050
			0.8140		4.7500	2.6300	2923	3.09566
					4.7200	2.7100	2906	2.60776
		ave	0.8115	1.0020	4.7267	2.6500		2.85797
	6	1	0.8180	1.0170	4.5300	2.9100	4116	2.56307
		2	0.8140	1.0120	4.4700	2.8900	4282	2.57353
3				4.4000	2.9900	5100	2.59456	
ave		0.8160	1.0145	4.5000	2.9300		2.57705	

Table A2 Interfacial tension of each phase in microemulsion formation with 0.1 wt% of Alfoterra at different NaCl concentrations and an initial oil to water ratio= 1:1

Alfoterra conc. (wt%)	NaCl conc. (wt%)	No.	Upper density (g/mL)	Lower density (g/mL)	Upper level	Lower level	Speed (rpm)	IFT (mN/m)
0.1	2	1	0.8000	0.9900	5.0400	2.2000	2223	0.2126
		2	0.8040	0.9820	4.9300	2.3200	2398	0.1921
		3			4.8700	2.4300	2724	0.2025
		ave	0.8020	0.9860	4.9467	2.3167		0.2024
	3	1	0.8320	1.0130	4.9000	2.2400	2085	0.1495
		2	0.8200	0.9970	4.8900	2.2600	2351	0.1837
		3			4.8300	2.2300	2367	0.1800
		ave	0.8260	1.0050	4.8733	2.2433		0.1666
	4	1	0.8110	1.0160	4.7000	2.5100	1784	0.0684
		2	0.8140	1.0100	4.6300	2.4800	2173	0.0961
		3			4.6100	2.5100	1988	0.0749
		ave	0.8125	1.0130	4.6650	2.5000		0.0798
	5	1	0.8030	1.0170	4.4600	2.6800	1710	0.0358
		2	0.8010	1.0120	4.3200	2.7800	1852	0.0272
		3			4.3100	2.7000	1752	0.0278
		ave	0.8020	1.0145	4.3633	2.7200		0.0302
	6	1	0.8000	1.0080	4.6800	2.9800	3742	0.1440
		2	0.8080	1.0100	4.6200	2.8100	3406	0.1440
3				4.5000	2.9400	3841	0.1172	
ave		0.8040	1.0090	4.6500	2.9100		0.1351	

Table A3 Interfacial tension of each phase in microemulsion formation with 0.1 wt% at different NaCl concentrations and an initial oil to water ratio = 1:1

Alfoterra conc. (wt%)	NaCl conc. (wt%)	No.	Upper density (g/mL)	Lower density (g/mL)	Upper level	Lower level	Speed (rpm)	IFT (mN/m)
0.15	2	1	0.8020	0.9840	4.8300	2.5900	3643	0.2711
		2	0.8100	0.9840	4.6900	2.6800	4208	0.2613
		3			4.7900	2.6500	3454	0.2125
		ave	0.8060	0.9840	4.7700	2.6400		0.2483
	3	1	0.8110	0.9930	5.0600	2.3500	2181	0.1803
		2	0.8080	0.9990	5.0600	2.3500	2181	0.1803
		3			4.9000	2.5200	3371	0.2917
		ave	0.8095	0.9960	5.0067	2.4067		0.2174
	4	1	0.8010	0.9970	4.6600	2.8100	3650	0.1688
		2	0.8020	0.9980	4.6900	2.7900	4093	0.2299
		3			4.6000	2.8200	3641	0.1496
		ave	0.8015	0.9975	4.6500	2.8067		0.1828
	5	1	0.8070	1.0040	4.5700	2.8800	2024	0.0400
		2	0.7990	0.9980	4.6100	2.8200	2609	0.0789
		3			4.5900	2.8700	2134	0.0468
		ave	0.8030	1.0010	4.5900	2.8567		0.0594
	6	1	0.7870	1.0020	4.5800	2.7800	2580	0.0836
		2	0.7970	1.0040	4.3900	2.9100	2819	0.0555
3				4.3600	2.8900	2856	0.0558	
ave		0.7920	1.0030	4.4433	2.8600		0.0696	

Table A4 Interfacial tension of each phase in microemulsion formation with 0.5 wt% at different NaCl concentrations and initial oil to water ratio = 1:1

Alfoterra conc. (wt%)	NaCl conc. (wt%)	No.	Upper density (g/mL)	Lower densi ty (g/mL)	Upper level	Lower level	Speed (rpm)	IFT (mN/m)
0.5	2	1	0.8160	0.9940	4.9100	2.3200	1787	0.0994
		2	0.8150	0.9880	4.9300	2.3100	1831	0.1080
		3			4.8000	2.4200	2182	0.1150
		ave	0.8155	0.9910	4.8800	2.3500		0.1075
	3	1	0.8060	0.9890	4.7100	2.6400	1704	0.0490
		2	0.8080	0.9980	4.6700	2.6900	1770	0.0463
		3			4.7000	2.6300	1703	
		ave	0.8070	0.9935	4.6933	2.6533		0.0477
	4	1	0.8240	1.0030	4.3400	3.0000	2811	0.0348
		2	0.8220	1.0020	4.3100	3.1000	3078	0.0308
		3			4.3200	3.0000	2944	0.0365
		ave	0.8230	1.0025	4.3233	3.0333		0.0340
	5	1	0.8230	1.0160	4.7500	2.6800	1911	0.0643
		2	0.8240	1.0200	4.6600	2.7200	2502	0.0908
		3			4.6500	2.8200	2478	0.0747
		ave	0.8235	1.0180	4.6867	2.7400		0.0775
	6	1	0.8260	1.0170				0.0000
		2	0.8110	1.0220	4.8100	2.5600	1857	0.0806
3				4.8000	2.5900	1906	0.0805	
ave		0.8185	1.0195	4.8100	2.5750		0.0805	

Appendix B Experimental Data of Colloidal Gas Aphron Studies.

1. CGA stability

The CGA stability was defined as the time that was required for the half volume of the initial liquid to drain

2. Gas hold up

The gas hold up was defined as the ratio of gas volume to the dispersion volume which was calculated by the following equations:

$$\text{Gas hold up} = \frac{V_g}{V_i} = \frac{V_i - V_l}{V_i} \quad (\text{B1})$$

where V_g = Gas volume
 V_l = Initial liquid volume
 V_i = Dispersion volume

3. Separation ratio

The separation ratio was defined as the ratio of oil concentration in aphron phase to the liquid phase which was calculated by the following equations:

$$\text{Separation ratio} = \frac{C_a}{C_l} \quad (\text{B2})$$

where C_a = concentration of oil in the aphron phase
 C_l = concentration of oil in the liquid phase

4. Enrichment Ratio

The enrichment was defined as the ratio of oil concentration in aphron phase to the initial solution which was calculated by the following equations:

$$\text{Enrichment ratio} = \frac{C_a}{C_i} \quad (\text{B3})$$

where C_a = concentration of oil in the aphron phase
 C_i = concentration of oil in the initial solution

5. Oil removal

The oil removal was calculated by the following equations:

$$\text{Oil removal (\%)} = \frac{C_i - C_l}{C_i} \quad (\text{B4})$$

where C_l = concentration of oil in the liquid phase
 C_i = concentration of oil in the initial solution

6. Effective parameter in colloidal gas aphron studies

6.1 Effect of stirring speed

Table B-1 Summary results for the effect of stirring speed in colloidal gas aphron studies at the system containing Alfoterra concentration = 0.1 wt%, NaCl concentration = 3 wt%, oil to water ratio = 1:19 at stirring time = 5 minute

System	Stirring speed (rpm)	CGA stability (sec)	Gas hold up	Separation ratio	Enrichment ratio	Oil removal (%)
1:19 Alf 0.1 N 3	4000	4.00	0.576	7.39	7.38	88.59
	5000	6.00	0.648	11.14	8.16	92.14
	6000	8.12	0.657	10.82	7.64	91.91
	7000	9.00	0.663	9.48	6.40	90.93
	8000	9.47	0.663	7.32	7.40	88.57

6.2 Effect of stirring time

Table B-2 Summary results for the effect of stirring time in colloidal gas aphron studies at the system containing Alfoterra concentration = 0.1 wt%, NaCl concentration = 3 wt%, oil to water ratio = 1:19 at stirring speed = 5000 rpm

System	Stirring time (min)	CGA stability (sec)	Gas hold up	Separation ratio	Enrichment ratio	Oil removal (%)
1:19 Alf 0.1 N 3	2	5.21	0.642	4.77	11.17	83.54
	5	6.00	0.648	11.14	8.16	92.14
	10	10.30	0.689	7.52	6.77	88.91
	15	17.04	0.694	5.38	6.69	84.98

6.3 Effect of surfactant concentration

Table B-3 Summary results for the effect of surfactant concentration in colloidal gas aphron studies at the system containing NaCl concentration = 3 wt%, oil to water ratio = 1:19 at stirring speed = 5000 rpm, stirring time = 5 minute

System	Surfactant conc. (wt%)	CGA stability (sec)	Gas hold up	Separation ratio	Enrichment ratio	Oil removal (%)
1:19 Alf 0.1 N 3	0.05	6.00	0.635	8.40	10.16	89.62
	0.1	6.00	0.648	11.14	8.16	92.14
	0.15	8.15	0.670	5.47	7.41	85.06
	0.5	16.27	0.703	4.83	9.43	83.25

6.4 Effect of NaCl concentration

Table B-4 Summary results for the effect of NaCl concentration in colloidal gas aphron studies at the system containing Alfoterra concentration = 0.1 wt%, oil to water ratio = 1:19 at stirring speed = 5000 rpm, stirring time = 5 minute

System	NaCl conc. (wt%)	CGA stability (sec)	Gas hold up	Separation ratio	Enrichment ratio	Oil removal (%)
1:19 Alf 0.1 N 3	2	10.25	0.689	11.27	5.57	91.57
	3	6.00	0.648	11.35	8.16	92.15
	4	4.46	0.635	6.99	10.3	87.51
	5	3.15	0.630	6.15	11.57	85.86
	6	2.48	0.612	4.88	12.85	82.67

Appendix C Experimental Data of Froth Flotation Experiment.

1. Dynamic oil removal

The oil removal was calculated by the following formulation:

$$\text{Oil removal (\%)} = \frac{(C_i - C_t) * 100}{C_i} \quad (\text{C1})$$

where C_t = concentration of oil in a solution at time t
 C_i = concentration of oil in a solution at time zero

2. Surfactant removal

The surfactant removal was calculated by the following formulation:

$$\text{Surfactant removal (\%)} = \frac{(C_{s,i} - C_{s,t}) * 100}{C_{s,i}} \quad (\text{C2})$$

where $C_{s,t}$ = concentration of surfactant in a solution at time t
 $C_{s,i}$ = concentration of surfactant in a solution at time zero

3. Enrichment Ratio

The enrichment ratio was calculated by the following formulation:

$$\text{Enrichment ratio} = \frac{C_f}{C_i} \quad (\text{C3})$$

where C_f = concentration of oil in the collapsed foam solution
 C_i = concentration of oil in the feed solution

4. Effective Parameter on froth flotation performance

4.1 Effect of air flow rate

Table C-1 Summary results for the effect of air flow rate on froth flotation performance at the system containing Alforterra concentration = 0.1 wt%, NaCl concentration = 3 wt% at oil to water ratio = 1:19

System	Air flow rate (l/min)	Oil removal (%)	Surfactant Removal (%)	Enrichment ratio	Foam flow rate (ml/min)	Foam wetness (g/ml)
1:19 Alf 0.1 N3	0.20	92.80	74.26	7.84	0.5	1.01
	0.25	91.52	70.54	7.11	0.52	1.02
	0.30	95.14	79.64	5.59	0.56	1.05
	0.35	91.50	79.41	4.92	0.55	1.00

4.2 Effect of colloidal gas aphon

Table C-2 Summary results for the effect of colloidal gas aphon on froth flotation performance at the system containing Alfoterra concentration = 0.1 wt%, NaCl concentration = 3 wt%, stirring speed = 5000 rpm, stirring time = 5 minute at oil to water ratio = 1:19

System	Air flow rate (l/min)	Oil removal (%)	Surfactant Removal (%)	Enrichment ratio	Foam flow rate (ml/min)	Foam wetness (g/ml)
1:19 Alf 0.1 N3 Speed 5000 Time 5	0.20	96.29	86.69	7.36	0.6	1.01
	0.25	96.11	86.91	6.46	0.75	1.06
	0.30	97.07	86.37	5.53	0.84	1.10
	0.35	95.51	83.41	4.58	0.76	1.06

4.3 Effect of equilibration time

Table C-3 Summary results for the effect of equilibration time on froth flotation performance at the system containing Alfoterra concentration = 0.1 wt%, NaCl concentration = 3 wt%, oil to water ratio = 1:19 at equilibration time of 1 month

System	Air flow rate (l/min)	Oil removal (%)	Surfactant Removal (%)	Enrichment ratio	Foam flow rate (ml/min)	Foam wetness (g/ml)
1:19 Alf 0.1 N3 Equilibrated	0.20	83.06	49.08	7.59	0.35	0.95
	0.25	83.08	52.59	5.26	0.35	0.98
	0.30	85.77	53.31	5.12	0.48	0.99
	0.35	83.63	57.07	4.61	0.41	0.96

Appendix D Experimental Data of Foam Ability and Foam Stability Experiment.

1. Foam ability

The foam ability was defined as the ratio of maximum foam height to initial solution height

$$\text{Foam ability} = \frac{H_{\max}}{H_i} \quad (\text{D1})$$

where H_{\max} := Maximum foam height
 H_i := Initial solution height

2. Foam Stability ($t_{1/2}$)

The foam stability was defined as the time that was required for the foam volume to collapse by half.

3. Effective parameter on foam ability and foam stability

3.1 Effect of air flow rate

Table D-1 Summary results for the effect of air flow rate on foam ability and foam stability at the system containing Alfoterra concentration = 0.1 wt%,

NaCl

concentration = 3 wt% at oil to water ratio = 1:19

System	Air flow rate (l/min)	Foam ability	Foam stability (min)
1:19 Alf 0.1 N3	0.20	9.09	2.17
	0.25	13.91	2.31
	0.30	20.27	3.25
	0.35	12.73	2.00

3.2 Effect of colloidal gas aphron

Table D-2 Summary results for the effect of colloidal gas aphron on foam ability and

foam stability at the system containing Alfoterra concentration = 0.1 wt%,

NaCl concentration = 3 wt% stirring speed = 5000 rpm, stirring time = 5

minute at oil to water ratio = 1:19

System	Air flow rate (l/min)	Foam ability	Foam stability (min)
1:19 Alf 0.1 N3 Speed 5000 Time 5	0.20	12.27	4.57
	0.25	17.45	4.20
	0.30	20.27	5.58
	0.35	20.27	5.04

3.3 Effect of equilibration time

Table D-3 Summary results for the effect of equilibration time on foam ability and foam stability at the system containing Alfoterra concentration = 0.1 wt%, NaCl = 3 wt%, oil to water ratio = 1:19 at equilibration time of 1 month

System	Air flow rate (l/min)	Foam ability	Foam stability (min)
1:19 Alf 0.1 N3 Equilibrated	0.20	1.18	0.51
	0.25	1.64	0.57
	0.30	2.45	1.05
	0.35	2.27	0.35

CURRICULUM VITAE

Name: Ms. Panita Angkathunyakul

Date of Birth: March 17, 1981

Nationality: Thai

University Education:

1999-2003 Bachelor Degree of Science in Chemical Engineering, Faculty of Science, Chulalongkorn University, Bangkok, Thailand